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Modelling chemistry in the nocturnal boundary layer above tropical rainforest and a generalised effective nocturnal ozone deposition velocity for sub-ppbv NO_x conditions

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Abstract

Measurements of atmospheric composition have been made over a re-
mote rainforest landscape. A box model has previously been demonstrated
to model the observed daytime chemistry well. However the box model
is unable to explain the nocturnal measurements of relatively high [NO]
and [O₃], but relatively low observed [NO₂]. It is shown that a one-
dimensional (1-D) column model with simple O₃-NO_x chemistry and a
simple representation of vertical transport is able to explain the observed
nocturnal concentrations and predict the likely vertical profiles of these
species in the nocturnal boundary layer (NBL). Concentrations of tracers
carried over from the end of the night can affect the atmospheric chem-
istry of the following day. To ascertain the anomaly introduced by using
the box model to represent the NBL, vertically-averaged NBL concentra-
tions at the end of the night are compared between the 1-D model and
the box model. It is found that, under low to medium [NO_x] conditions

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($\text{NO}_x < 1$ ppbv), a simple parametrisation can be used to modify the box model deposition velocity of ozone, in order to achieve good agreement between the box and 1-D models for these end-of-night concentrations of NO_x and O_3 . This parametrisation would also be used in global climate-chemistry models with limited vertical resolution near the surface. Box-model results for the following day differ substantially if this effective nocturnal deposition velocity for ozone is implemented; for instance, there is a 9% increase in the following days peak ozone concentration. However under medium to high [NO_x] conditions ($\text{NO}_x > 1$ ppbv), the effect on the chemistry due to the vertical distribution of the species means no box model can adequately represent chemistry in the NBL without modifying reaction rate coefficients.

1 Introduction

Box modelling studies are frequently used to investigate daytime chemistry in the planetary boundary layer (PBL) (e.g. Aumont et al, 2003; Toyota et al, 2004; Hamer and Shallcross, 2007; Read et al, 2008; Pugh et al, 2010) because the computational expense can be focused on solving the chemistry rather than solving the atmospheric transport. Box models are appropriate during most daytime conditions (assuming the modelling domain is horizontally homogeneous on relevant advective timescales), as the rapid mixing within the PBL means that their critical assumption of instantaneous mixing throughout the model box is valid. As a result, box model output can be compared with point measurements made in the PBL during the day. In contrast, the nocturnal boundary layer (NBL) over land often assumes different characteristics. The lack of solar heating at the ground surface results in less energy in the system, and typically a lower boundary layer height than during the day. Often, particularly in the absence of a strong wind to induce mechanical turbulence, this lack of thermally-induced turbulence leads to a statically stable NBL, with very limited mixing. The thermal effects are particularly important in the tropics where insolation is high and synoptic wind speeds are often low. Accordingly, under such conditions, very considerable gradients of species may exist in the NBL, meaning that point measurements may not be suitable for comparison with a box model representing the average NBL concentrations of the measured species. Furthermore, the ability of the box model to represent chemical and physical processes within the NBL may be compromised under statically stable conditions (e.g. Geyer and Stutz, 2004a; Galmarini et al, 1997).

This study compares results from the chemistry box model CiTTyCAT (Wild et al, 1996; Evans et al, 2000; Emmerson et al, 2004; Donovan et al, 2005; Real et al, 2007, 2008; Hewitt et al, 2009; Pugh et al, 2010) to measurements made above a remote tropical rainforest, and finds that the box model is unable to explain the unexpectedly high night-time concentrations of NO. A previous study with CiTTyCAT (Pugh et al, 2010) demonstrated that the box model could reproduce day-time chemistry, including OH and peroxy radical observations, reasonably well, particularly if inhomogeneous mixing (sometimes called segregation) was parametrised using an effective rate coefficient for the isoprene + OH reaction. First some background information on nocturnal NO_x chemistry and modelling is reviewed in Section 2. The measurements are described in Section 3 and compared to the box model in Section 4, and reasons for the possible discrepancies are discussed. Section 5 demonstrates that a one-dimensional (1-D) column model can replicate the main features of the night-time measurements. Average

NBL concentrations in the 1-D model are then compared with those in the box model to assess the extent to which the box model output deviates from that provided by the 1-D model. Section 6 then seeks to test whether the box model can be adapted to provide an adequate representation of the chemical evolution of the NBL, where adequate means, in this instance, sufficient to confirm known chemistry and sufficient to have a minimal impact on later daytime chemical evolution. A summary follows in Section 7.

2 NO_x and O_3 in the NBL

Nocturnal NO_x ($\text{NO} + \text{NO}_2$) partitioning in the remote boundary layer is typically strongly skewed in favour of NO_2 , in the absence of very high emission rates of NO . Very little NO is present due to its loss via $\text{NO} + \text{O}_3 \rightarrow \text{NO}_2 + \text{O}_2$ (R1) and the absence of NO_2 photolysis via $\text{NO}_2 + h\nu \rightarrow \text{NO} + \text{O}(^3\text{P})$ (R2). In most situations above land, there is a surface emission of NO , which can lead to a depletion of ozone in the nocturnal boundary layer (NBL) and an accumulation of NO_2 over the course of the night. Other potentially important reactions are the formation of NO_3 via $\text{NO}_2 + \text{O}_3 \rightarrow \text{NO}_3 + \text{O}_2$ (R3) and of N_2O_5 via $\text{N}_2\text{O}_5 + \text{M} \rightleftharpoons \text{NO}_2 + \text{NO}_3 + \text{M}$ (R4). Both NO_3 and N_2O_5 are unstable when exposed to photolytic flux, but during the night can act as stores or sinks (should they be deposited) of NO_x . N_2O_5 can be lost due to a heterogeneous reaction with water on the surface of aerosol particles via $\text{N}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$ (R5). NO_3 is lost by reaction with a range of species, but notably by $\text{NO}_3 + \text{NO}_2 \rightarrow \text{NO} + \text{NO}_2 + \text{O}_2$ (R6). A comprehensive review of night-time chemistry is given by Wayne et al (1991) or Monks (2005). Dry deposition is also an important night-time sink for many species, particularly NO_2 and O_3 . However deposition rates over vegetation, particularly for ozone, are often lower at night-time than during the day, as plant stomata tend to close at night. Wet deposition is not considered here as night-time rainfall was limited during the study period (Hewitt et al, 2010). For a more comprehensive review of biosphere-atmosphere exchange of reactive compounds the reader is referred to Fowler et al (2009).

Previous studies have shown success in representing night-time chemistry using one dimensional models. Hov (1983) investigated PBL chemistry using a 1-D K-profile model and demonstrated that many species, particularly NO_x and O_3 , should be expected to show marked vertical gradients in a stable NBL. NO in particular displays a high concentration layer close to its surface source under these conditions, whereas $[\text{O}_3]$ typically has a minima at the ground due to dry deposition and NO titration, and features a positive gradient throughout the NBL. This typical pattern has been confirmed by several measurement studies, including Cros et al (1992), Pisano et al (1997), Gusten et al (1998), Glaser et al (2003), Stutz et al (2004), Geyer and Stutz (2004a) and Brown et al (2007).

Fitzjarrald and Lenschow (1983) and Gao and Wesely (1994) have shown that fluxes of reactive species, particularly NO , NO_2 , NO_3 and N_2O_5 , can be modified by chemical reactions, such that the surface NO flux can be reduced by 80% at a height of 100 m, even during daytime conditions. Such non-linear effects of chemistry on fluxes are not captured by K-profile models, and are usually solved using a second-order closure model. Geyer and Stutz (2004a) developed the modified K model to take account of the effects of chemistry on flux. They applied their model to the case of NO_x chemistry in the urban and rural NBL, finding that dry deposition often led to positive vertical gradients of O_3 and NO_2 . The conclusions derived from their analysis

were found to be in good agreement with field measurements of vertical profiles of NO, NO₂ and O₃ by Stutz et al (2004), suggesting that such models can effectively represent NBL species profiles. Similar vertical profiles of NO₂ and O₃ were observed by Brown et al (2007). Gao and Wesely (1994) found that NO₂ deposition at the surface could be enhanced due to an increase in [NO₂] near the surface NO source.

Galmarini et al (1997) used a 1-D model to study the evolution of NO_x in the NBL, under moderate NO_x mixing ratios of a few ppbv. They found substantial departures between a box model and the vertically averaged concentrations from the 1-D model for NO, NO₂, NO₃, N₂O₅ and HNO₃, although O₃ showed little deviation. Galmarini et al (1997) show that they can force the box model to accurately calculate the average NBL concentrations by utilising an effective reaction rate coefficient incorporating a term representing the subgrid effects. However the magnitude of the subgrid term will vary with the situation, requiring the effective rate coefficients to be revised for each new application using a 1-D model. Vinuesa and Vilà-Guerau de Arellano (2005) describe a similar method for the convective boundary layer.

To date, we know of no method which can improve the night-time performance of box models without requiring calibration for each case study (cf. Galmarini et al, 1997). In the analysis that follows, using a modified K model, it is shown that the vertical variations in night-time concentrations of O₃ and NO_x expected from the above-mentioned studies are consistent with measurements made at the OP3 rainforest site. A generalised parametrisation is then developed, which may be used to improve model estimation of daybreak concentrations of O₃ for any site with similar low [NO_x] characteristics. This parametrisation offers a method to incorporate the net result some of the complexities of NBL processes into a box (or indeed global) model, without requiring recalibration for each case study.

3 Measurement description

The OP3-Danum-08 field campaign was carried out in three phases during 2008: (i) ground-based measurements only, during April/May, (ii) a reduced set of flux and concentration measurements at a nearby oil palm estate during May/June, and (iii) both ground and aircraft measurements, during June/July. We focus here on measurements made during the first campaign, OP3-1.

A comprehensive suite of ground-based measurements of atmospheric composition were made at the Bukit Atur Global Atmosphere Watch station (4°58'59" N, 117°50'39" E) (Hewitt et al, 2010). This measurement site was situated at 437 m a.m.s.l., in a small (0.75 ha) grassy clearing, on a hill rising ~260 m above the surrounding valley floor. The clearing is surrounded by secondary rainforest, rising to ~10 m above the hilltop on three sides. A 100 m high, steel framed measurement tower was set in the centre of the clearing. A detailed description of the site is given in Hewitt et al (2010). This study concentrates on the measurements of NO, NO₂ and O₃ made at 5 m above ground level at Bukit Atur. NO was measured directly by the chemiluminescence technique, NO₂ was measured by conversion to NO via UV photolysis and subsequent detection of the NO by chemiluminescence and O₃ was measured using a Thermo Electron Instrument (TEI) Model 49i UV Analyser. A detailed description of the measurements is given by Pike et al (2009). All three measurements were co-located. In addition, during OP3-3, gradient measurements of O₃ were made at heights of 30, 45, 60 and 75 m on the measurement tower using a TEI Model 49C UV photometric ozone analyser, with a precision of 1 ppbv.

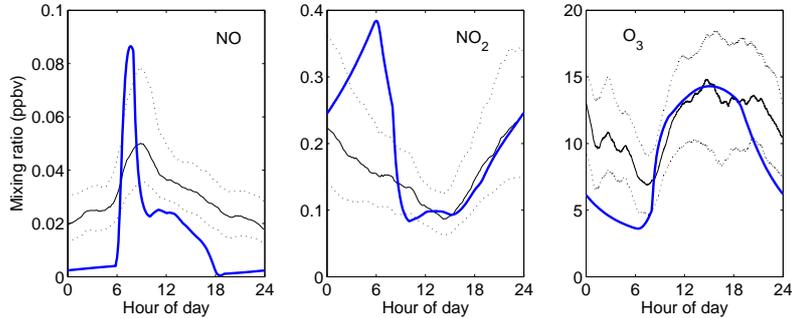


Figure 1: Solid black lines show median NO (left panel), NO₂ (centre panel) and O₃ (right panel) mixing ratios, measured at 5 m above ground at Bukit Atur during OP3-1, as a function of local time. Dotted lines show the upper and lower quartiles. The solid blue lines show the results of the standard CiTTyCAT run.

Measurements of night-time NO made at 5 m during the OP3 campaigns revealed substantial concentrations persisting throughout the night. Figure 1 shows that median measurements for OP3-1 give a fairly constant background NO mixing ratio of ~ 20 pptv. This is in contrast to previous studies of rural night-time chemistry, which have found very low mixing ratios of NO, of the order of a few pptv (Biesenthal et al, 1998; Faloon et al, 2001; Mihele and Hastie, 2003). Median night-time NO₂ mixing ratios remain relatively low, compared to these studies, at around 200 pptv.

4 Box Modelling

4.1 Box model set-up

The CiTTyCAT model was run in two-box mode as described in Pugh et al (2010). Between 0800 and 1800LT, when solar heating means the well-mixed assumption may be considered valid for the whole PBL, a single box is integrated over the height of the PBL (indicated to be 800 m by Dopplar LIDAR measurements, Pearson et al, 2010). However at 1800LT (sunset) the single box is split into two, with a lower box, representing the NBL, integrated up to 200 m, and an upper box, representing the residual layer, integrated between 200 m and 800 m. No mixing between these layers is allowed during the night; however, when turbulent mixing is re-initiated between 0800-1000LT the following morning (Pearson et al, 2010), the mixing height rises linearly to 800 m and the residual layer concentrations are gradually mixed into the lower box until the lower box has engulfed the residual layer. For details of the mixing parametrisation see Pugh et al (2010). No mixing vertically, out of the upper box, or laterally is permitted. This two layer model allows explicit consideration of the result of restricting the effect of emissions and deposition to a lower volume of atmosphere (the NBL) during the night. In the absence of appropriate measurements, the 200 m lower box height was selected by Pugh et al (2010) by comparing the results of model sensitivity studies with measured daytime concentrations. However, as the results of this study concentrate on comparisons between models, the most important factor is

that a consistent height is chosen for all analysis.

Emissions of NO, isoprene and monoterpenes were prescribed following Pugh et al (2010) with 24-hour average fluxes corresponding to 5.6×10^9 and, 5.73×10^{10} and 1.8×10^{10} molecules $\text{cm}^{-2} \text{s}^{-1}$ respectively. Isoprene and monoterpene emissions follow a principally light-dependent cycle defined by the MEGAN algorithm (Guenther et al, 2006), whereas NO emissions are assumed constant. Dry deposition is included for O_3 , NO_x , N_2O_5 , HNO_3 , HNO_4 , H_2O_2 , HCHO , CH_3CHO , CH_3OOH , $\text{CH}_3\text{CO}_3\text{H}$, PANs and isoprene nitrates using the values listed in Evans et al (2000). The deposition velocity for isoprene nitrates is assumed to be as for HNO_3 , following Horowitz et al (2007). A sinusoidal diurnal variation is imposed as in Evans et al (2000), by multiplying the prescribed deposition velocity by a factor s_d , where

$$s_d = 0.5 + (0.5 \times \sin(15 \times (t_d - 6))), \quad (1)$$

and t_d is the time of day expressed in hours. This leads to an average night-time $V_{d(\text{O}_3)} = 0.6 \text{ cm s}^{-1}$ over Bukit Atur. Wet deposition is applied for soluble species as in Pugh et al (2010).

4.2 Comparison of box model with measurements

Figure 1 shows the CiTTYCAT lower box concentrations as compared to the OP3-1 median measurement. A good agreement between model and measurements is achieved for the well-mixed period of the daytime (1000-1800LT) for NO_2 , NO and O_3 . Night-time concentrations show very poor agreement; modelled night-time [NO] is $\sim 7 \times$ smaller than observations and night-time NO_2 accumulates throughout the night to mixing ratios of over 500 pptv. The overestimation of modelled [NO] at sunrise is due to the photolysis of the overestimated night-time [NO_2] accumulation (R2). After 0800 LT, mixing in of low- NO_2 air from the upper box rapidly decreases the NO_2 mixing ratio in the lower box, eliminating the NO peak by 1000 LT.

There are two hypotheses that could explain these measurements, either (1) underestimation of the NO emission, or (2) limited mixing within the NBL causing the restriction of the high NO concentrations to a small region of the NBL, meaning that they are not representative of the average NBL concentration modelled by CiTTYCAT.

Hypothesis (1) is investigated by running CiTTYCAT with NO concentrations constrained to those measured. This leads to a huge build-up in night-time NO_2 mixing ratios, in excess of 2 ppbv, demonstrating that simply increasing emissions cannot improve modelling of night-time [NO_x]. However, it is possible that some process, perhaps hitherto unrepresented in CiTTYCAT, could remove this excess NO_2 . Running CiTTYCAT with NO concentrations constrained to those measured and the NO_2 deposition velocity set to the upper end of the range reported in the literature (0.8 cm s^{-1}), only leads to a 1 ppbv reduction in peak night-time NO_2 to 1.7 ppbv. Setting NO_2 deposition velocity equal to that used for HNO_3 (2.0 cm s^{-1}) still results in peak [NO₂] four times the measurement. Therefore we rule out poor representation of NO_2 deposition as a likely cause. We note that using a lower NBL height in the model would modify simultaneously both deposition velocities, and the volume throughout which emissions are diluted. Running CiTTYCAT with a lower box height reduced by 50% to 100 m, improves the model fit to [NO], but, as above, overestimates [NO₂], whilst also worsening the fit for [O₃] by increasing O₃ deposition.

Alternatively, the excess NO_2 might be being sequestered chemically. Investigating all the reactions in CiTTYCAT capable of sequestering NO_2 (See Supplementary

Information) shows that either an increase in the rate coefficient by a magnitude much greater than the IUPAC reported uncertainty, or an increase in the concentration of the other reactant well beyond the measured concentration is required. Kleffmann et al (1998) and Kleffmann (2007) describe methods of HONO formation via the heterogeneous reaction of NO₂ with either water or hydrocarbons. However measurements of HONO during OP3-3 indicated mixing ratios of less than 10 pptv (Pugh et al, 2010); much too small for HONO formation to provide the necessary NO₂ sink. It is possible that some unknown process may be able to convert NO₂ to NO in the absence of sunlight. NO will react with peroxy radicals, whose mixing ratios average 10 pptv during the night (Pugh et al, 2010). This will lead to OH formation, as observed in the night-time OH concentrations of 5×10^5 molecules cm⁻³. Therefore the measurements of OH and peroxy radicals made at 5 m are consistent with this hypothesis. However, if this process were occurring throughout the NBL, it would be expected that hydrocarbon concentrations would show signs of oxidation occurring throughout the NBL. Isoprene concentrations at 5 m, decrease on average throughout the night (C. Jones, pers. comm.), but isoprene concentrations at 75 m are much more stable, particularly in the period 0000 to 0600 LT, and also much higher than at 5 m. This suggests that the conditions of night-time OH production, and hence the elevated NO concentrations, do not prevail throughout the boundary layer. Indeed the difference in isoprene concentrations at 5 m and 75 m suggests weak NBL mixing, supporting hypothesis (2).

Limited mixing within the NBL (hypothesis 2) could arise due to stable stratification of the NBL suppressing mixing during the night, leading to high [NO] near its surface source, as discussed in Section 2. Hypothesis (2) is tested in the next section using a 1-D model.

5 1-D modelling

5.1 1-D model description and setup

A gas-phase 1-D column model, based on MANIC (Microphysical Aerosol Numerical model Incorporating Chemistry) (described in Lowe et al, 2009; Ryder et al, 2010), has been used. This model is constructed using the Kinetic PreProcessor (KPP) (Damian et al, 2002; Sandu et al, 2003; Daescu et al, 2003). In this work the aerosol chemistry and microphysics features are not implemented, nor any of the in-canopy effects. Operator splitting is used to perform the chemistry and deposition procedures for each level in the column for one timestep, followed by the vertical transport for each species in the column. The approach is similar to earlier chemistry and transport models such as Gao et al (1993) and (ApSimon et al, 1994). The model consists of 58 boxes logarithmically distributed between 0 and 200 m, according to the equation,

$$z_{top} = exp\left(\frac{l + 13}{11.7}\right) - 3 \quad (2)$$

where z_{top} is the top of the level and l is the level number. A simple chemistry scheme limited to NO_x and O₃ chemistry, is integrated within each box (Table 1). No photochemistry is included as all integrations in this study are performed at night. Radical chemistry, such as initiated by oxidation of terpenes, is neglected for simplicity. Deposition from the lowermost box is included for O₃, NO₂ and HNO₃ at deposition

velocities of 0.6, 0.3 and 2.3 cm s^{-1} respectively¹, these being the average night-time deposition velocities used in CiTTYCAT following Evans et al (2000). Mixing between the boxes is parameterised using K-theory,

$$j_i(z, t) = -K(z, t) \frac{\partial c_i(z, t)}{\partial z} \quad (3)$$

where K is a diffusion coefficient and j_i is the vertical flux of a gas, i , whose concentration is given by c_i . Equation 3 is only valid for inert or slowly reacting gases because the effect of chemistry on the vertical flux of the gas is not considered (see e.g. Fitzjarrald and Lenschow, 1983; Brost et al, 1988; Vila-Guerau de Arellano, 2003). Geyer and Stutz (2004a) suggest a modified K model in which,

$$k_i(z, t) = K(z) \left(1 + \frac{\frac{\partial(P_i(z, t) - L_i(z, t))}{\partial z} \tau_{ieff}(z, t)}{\frac{\partial c_i(z, t)}{\partial z}} \right) \quad (4)$$

where P_i and L_i are chemical production and loss respectively and τ_{ieff} is the effective timescale as approximated by Hamba (1987). Geyer and Stutz (2004a) found that it was sufficient only to apply the modified K model for NO, and that approach has been taken here. K was approximated via Monin-Obukhov similarity theory following Businger et al (1971) using,

$$K(z, t) = \frac{\kappa u^*(t) z}{\Phi_s \left(\frac{z}{L}, t \right)} \quad (5)$$

where κ is the von Karman constant (0.38) and u^* the friction velocity. The correction factor Φ for a stable atmosphere is defined by Seinfeld and Pandis (2006) as,

$$\Phi \left(\frac{z}{L}, t \right) = 1 + 4.7 \frac{z(t)}{L(t)} \quad (6)$$

where L is the Monin-Obukhov length. As OP3-1 campaign median measurements for L and u^* do not show a notable trend though the night, a single profile of K is used during this period. The calculated K profile is shown in Fig. 2, and is of a similar magnitude to, although slightly smaller than, that used by Geyer and Stutz (2004a). The lower quartile shows that there is considerable variation in this number throughout the night. The upper quartile is not shown as its value above ~ 10 m is $\gg 1.0 \text{m}^2 \text{s}^{-1}$, indicating that some relatively rapid transport events do occur during the night. These rapid transport events, which would cause short-lived perturbations to the steady-state profile, are not considered further in this analysis.

5.2 Comparison of 1-D model with measurements

Figure 3 shows the results of the 1-D model for the OP3-1 scenario, for 12 hours between 1800 and 0600LT, using the same inputs and initial conditions as CiTTYCAT. As expected, modelled NO concentrations are high near the ground, but tail-off quickly to very small (less than 1 pptv) concentrations above ~ 30 m. Conversely modelled $[\text{O}_3]$ is depleted near the ground. Ozone loss near the surface could be either due to reaction with freshly-emitted NO (R1), or dry deposition. However the total number of NO molecules emitted overnight in the model ($2.4 \times 10^{14} \text{ cm}^{-2}$) is only 4% of the number of ozone molecules present in the NBL at sunset ($6.0 \times 10^{15} \text{ cm}^{-2}$), demonstrating

¹In comparisons between the 1-D model and CiTTYCAT in section 6, these same deposition velocities are used in CiTTYCAT.

Table 1: Chemical scheme for the 1-D model. Reaction rates for termolecular reactions are calculated using $k = \left(\frac{k_0[M]}{1+k_0[M]/k_\infty} \right) F^{(1+[\log_{10}(k_0[M]/k_\infty)]^2)^{-1}}$ where M is the air density in molecules cm^{-3} . R is the universal gas constant and T is temperature in Kelvin.

Reactants	Products	Rate coefficient ($\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$)
$\text{NO} + \text{O}_3$	$\rightarrow \text{NO}_2 + \text{O}_2$	$3.0 \times 10^{-12} \exp(-1500/RT)$
$\text{NO} + \text{NO}_3$	$\rightarrow 2 \text{NO}_2$	$1.5 \times 10^{-11} \exp(170/RT)$
$\text{NO}_2 + \text{O}_3$	$\rightarrow \text{NO}_3 + \text{O}_2$	$1.2 \times 10^{-13} \exp(-2450/RT)$
$\text{NO}_2 + \text{NO}_3 + \text{M}$	$\rightarrow \text{N}_2\text{O}_5 + \text{M}$	$k_0 = 2.0 \times 10^{-30} (T/300)^{-4.4}$ $k_\infty = 1.4 \times 10^{-12} (T/300)^{0.7}$ $F = 0.6$
$\text{N}_2\text{O}_5 + \text{M}$	$\rightarrow \text{NO}_2 + \text{NO}_3$	Coefficient of $\text{NO}_2 + \text{NO}_3 + \text{M}$ $/ 2.7 \times 10^{-27} \exp(11000/RT)$
$\text{NO}_3 + \text{NO}_3$	$\rightarrow 2 \text{NO}_2 + \text{O}_2$	$8.5 \times 10^{-13} \exp(-2450/RT)$
$\text{N}_2\text{O}_5 + \text{H}_2\text{O}$	$\rightarrow 2 \text{HNO}_3$	2.6×10^{-22}
$\text{CO} + \text{NO}_3$	$\rightarrow \text{NO}_2 + \text{CO}_2$	$3.03 \times 10^{-12} \exp(-446/RT)$
$\text{CO} + \text{O}_3$	$\rightarrow \text{O}_2 + \text{CO}_2$	$7.86 \times 10^{-15} \exp(-1913/RT)$

that in this scenario R1 is a relatively minor ozone sink compared to dry deposition. This conclusion is confirmed by the sensitivity studies carried out in this section (e.g. Fig. 5). The modelled gradient in ozone agrees well, in terms of shape, with gradient measurements made at the site during OP3-3 (black dots in Fig. 3). Modelling is not carried out here for OP3-3 because a localised NO leak, from one of the measurement instruments, makes the NO_x measurements challenging to interpret in this context.

Modelled NO_2 concentrations show a weak positive gradient near the ground, with a maximum between 3 m and 9 m, followed by a clear negative gradient above this point. As NO_2 is principally formed by reaction R1, it follows that NO_2 is nearly all formed in the near-surface layer, but is then mixed upwards. Upward mixing prevents the build-up of high concentrations of NO_2 seen when CiTTyCAT is constrained to the 5-m [NO] measurements. Figure 4 shows that the 1-D model is able to maintain a high 5-m NO concentration throughout the night without producing an erroneously high NO_2 concentration. The net flux of NO_x through reaction (R4), taking into account the counteracting effect of the reverse reaction, is 0.7% of the NO_x flux through R1, showing that reaction R4 is of minor significance at these relatively low (in global terms) concentrations of NO_x . The night-time and vertically averaged mixing ratios of NO_3 and N_2O_5 are 0.1 pptv and 0.03 pptv respectively.

Although the 1-D model results above support the stratification hypothesis described in section 4.2, Fig. 4 shows that the 1-D model calculations do not match the magnitude of the OP3-1 measurements: modelled NO and NO_2 concentrations at 5 m are too high, and the O_3 concentration is too low. There are several factors which may be responsible for this disagreement, including the rate of mixing, emission and deposition. These factors are investigated through sensitivity studies, which are summarised below. The impact of the reduced chemical complexity in the 1-D model, as compared to CiTTyCAT, is found to be not important for this study (see

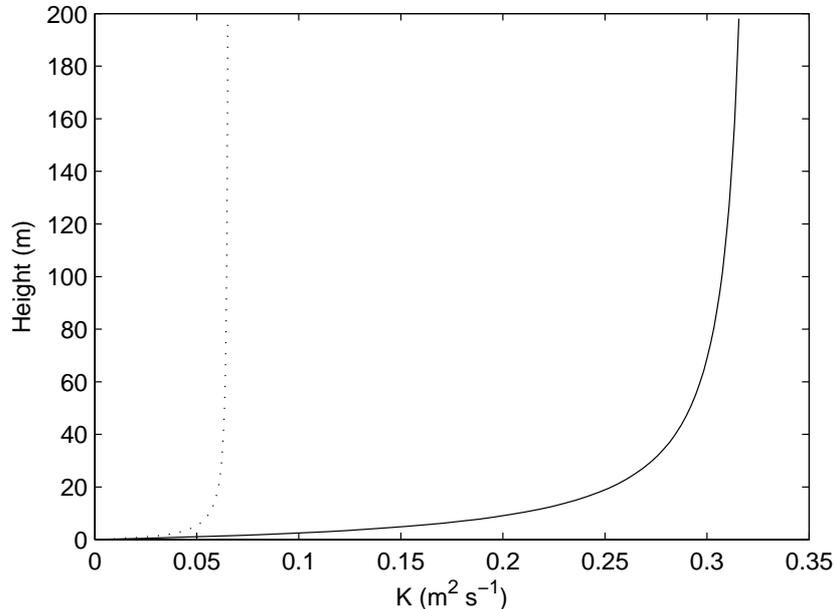


Figure 2: OP3-1 median profile of the diffusion coefficient, K , with height averaged throughout the night. The dotted line shows the low quartile of the night-time average.

Supplementary Information).

The model-measurement agreement for NO_x and O_3 is largely insensitive to the K -profile used. Doubling K results in a slightly more uniform tracer profile with height, whilst halving K accentuates the peaks. However, the profiles remain similar to those in Fig. 3. The NO_2 deposition rate also has a limited effect on the model-measurement agreement. In order to bring modelled 5-m $[\text{NO}_2]$ into agreement with the measurements an increase in the NO_2 deposition velocity from 0.3 cm s^{-1} to 4.0 cm s^{-1} is required. In the context of literature values for night-time NO_2 deposition velocity of $0.1 - 0.8 \text{ cm s}^{-1}$ (Wesely and Hicks, 2000; Watt et al, 2004; Michou et al, 2005; Trebs et al, 2006; Zhang et al, 2009), increasing the NO_2 deposition rate is not an appropriate method with which to improve the model fit to 5-m $[\text{NO}_2]$ measurements. In contrast, reducing the model ozone deposition velocity to the lower end of the literature range of $0.1 - 0.6 \text{ cm s}^{-1}$ (Fan et al, 1990; Wesely and Hicks, 2000; Matsuda et al, 2005; Michou et al, 2005; Coyle et al, 2006; Kerckweg et al, 2006; Rummel et al, 2007), brings 5-m modelled concentrations of O_3 and NO into excellent agreement with the observations.

The NO emission rate used above of $5.6 \times 10^9 \text{ molecules cm}^{-2} \text{ s}^{-1}$, estimated using CiTTyCAT by Pugh et al (2010), was optimised using 1000-1800 LT measurements, when the boundary layer could be considered well-mixed. In a well-mixed boundary layer, concentrations at any one point would be influenced by a relatively large emission footprint. In contrast, the NO measurements made at 5 m in the stratified night-time

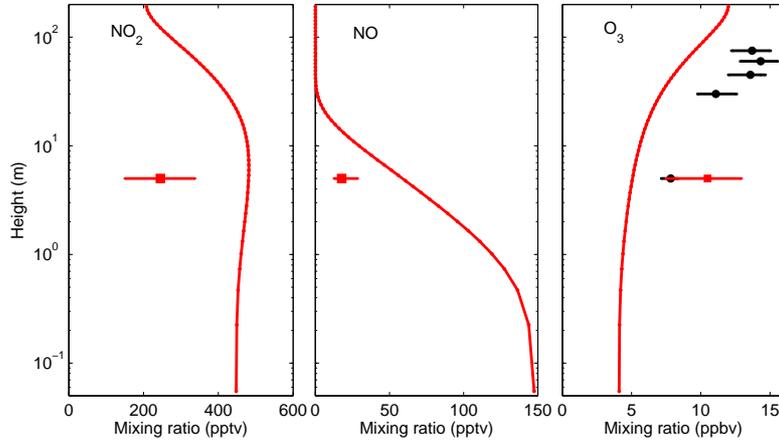


Figure 3: 1-D model vertical profile at midnight (red line, each level output marked by dots) and OP3-1 median 5-m measurements at that time (red squares, upper and lower quartiles shown by bars). Gradient measurements of median ozone concentration made during OP3-3 are shown (black dots, upper and lower quartiles shown by bars) to compare the profile shapes (note that the magnitude of OP3-3 concentrations is not directly comparable to those from OP3-1).

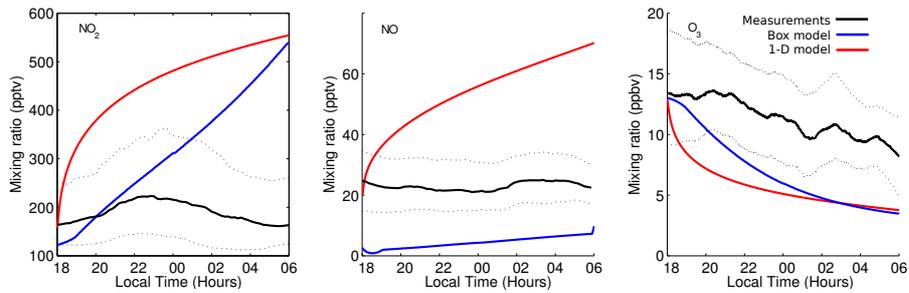


Figure 4: Comparison of 1-D model output at 5 m throughout the night (red line) against OP3-1 median 5-m measurements (black line, upper and lower quartiles shown by dots). CiTTyCAT lower-box output over the same time is shown by the blue line.

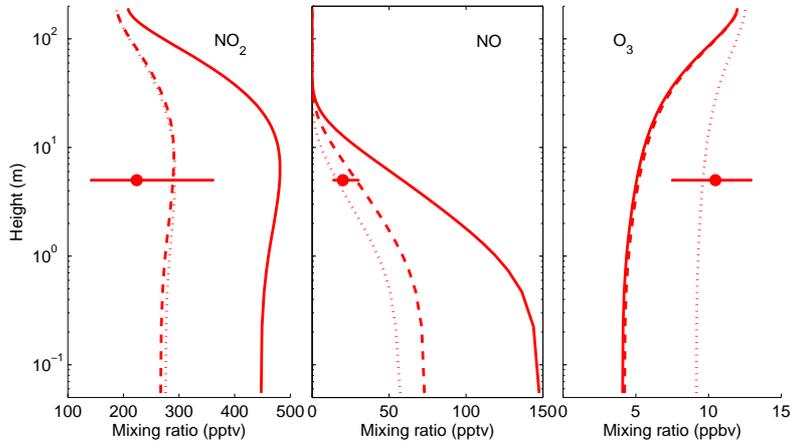


Figure 5: Comparison of measurements (large dot) against model 6 hour (mid-night) concentrations with standard NO emission (solid line), NO emission reduced by 50% (dashed line), and NO emission reduced by 50% and $V_{d(O_3)}=0.1 \text{ cm s}^{-1}$ (dotted line).

boundary layer at Bukit Atur are likely to have a much smaller emission footprint, most likely with a different NO emission potential as a result of local heterogeneities. Furthermore, diurnal temperature variations could result in a smaller NO flux at night (Yienger and Levy, 1995). Figures 5 and 6 show the effect of reducing the NO emission by 50%, which brings model and measurements much more into line with regard to NO and NO_2 concentrations. Combining the NO emission decrease with an O_3 deposition velocity of 0.1 cm s^{-1} generates a model output within variability of measurements of NO, NO_2 and O_3 throughout most of the night.

Intermittent nocturnal mixing events (e.g. Fitzjarrald and Moore, 1990; Acevedo et al, 2006) provide an alternative explanation to the model-measurement discrepancy for NO_x and O_3 , with the residual layer most likely providing a sink for NBL NO_x and a source for NBL O_3 . An increase in nocturnal mixing events could also explain the turnover in $[\text{NO}_2]$ after midnight. However, the NBL height used here was optimised by Pugh et al (2010) using measured tracer concentrations. Therefore, the average effect of such mixing events have effectively been incorporated into their calculation of the 200 m NBL height. This approach was also adopted by Strong et al (2009). Nonetheless, despite the failure of the simple 1-D model used here to capture all the complexities of the situation, the results demonstrate that the measured night-time concentrations at 5 m can be explained by the stratification hypothesis and therefore it is critical to consider stratification and limited mixing when comparing with night-time measurements at this site.

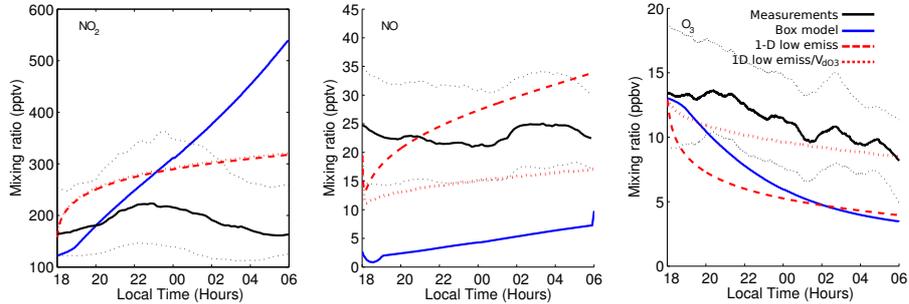


Figure 6: Timeseries comparison of 5-m OP3-1 median measurements (black solid line, upper and lower quartiles shown by black dotted line) against 1D model 5-m concentrations with NO emission reduced by 50% (red dashed line) and NO emission reduced by 50% and $V_{d(O_3)}=0.1 \text{ cm s}^{-1}$ (red dotted line). CiTTYCAT lower-box output over the same time is shown by the blue line.

6 Comparing box and 1-D models in the NBL

A 1-D model is able to produce a better fit to the measured nocturnal concentrations of NO_x and O_3 at 5 m, compared to the CiTTYCAT box model. However, as box models are typically used for modelling daytime tracer concentrations, a pertinent question is how well can CiTTYCAT calculate vertically-averaged NBL concentrations at the end of the night? As the end-of-night concentrations effectively provide the initial conditions for the chemistry during the day, failing to model these daybreak concentrations is likely to result in inaccuracies in daytime chemistry, particularly in the morning. To address this question CiTTYCAT and the 1-D model were run with a NBL height of 200 m. Emissions of NO and deposition velocities were equalised between the two models to the average night-time values used in the CiTTYCAT run in Section 4 of $5.6 \times 10^9 \text{ molecules NO cm}^{-2} \text{ s}^{-1}$, $V_{d\text{NO}_2}=0.3 \text{ cm s}^{-1}$, $V_{d\text{O}_3}=0.6 \text{ cm s}^{-1}$ and $V_{d\text{HNO}_3}=2.3 \text{ cm s}^{-1}$. Note that the emission and deposition rates from Section 5.2 are not used here, as the intention is to test the daytime-optimised CiTTYCAT setup. Both models were initialised with average measured 1800 LT concentrations (see Supplementary Information) and run for 12 hours until daybreak.

It has been shown (e.g. Hov, 1983; Fitzjarrald and Lenschow, 1983; Thompson and Lenschow, 1984; Gao and Wesely, 1994; Geyer and Stutz, 2004a), that in the presence of a NO concentration substantially greater than that observed during OP3, the titration of O_3 by NO (R1) can have very large effects on O_3 and NO_2 concentrations, whilst also preventing enough O_3 and NO_2 reaching the surface to generate a notable deposition flux. In order to understand whether the results produced by this model comparison are limited to the low $[\text{NO}_x]$ conditions encountered during OP3, the comparison is also carried for medium and high $[\text{NO}_x]$ conditions, as might be encountered at other locations.

6.1 Low $[\text{NO}_x]$ conditions

The average integrated NBL concentrations at daybreak (0600 LT) are compared in Fig. 7. NO and NO_2 mixing ratios show very close agreement. The close inter-model

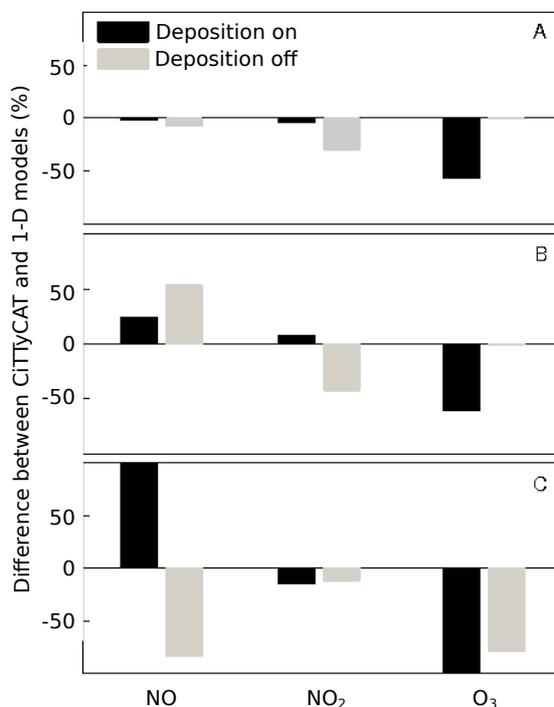


Figure 7: Percentage difference between CiTTYCAT and the 1-D model for vertically averaged NBL concentrations at daybreak $\left(\frac{\text{CiTTYCAT} - 1\text{D}}{1\text{D}} \times 100\right)$. Deposition on uses dry deposition velocities as specified at the start of this section. Deposition off sets all dry deposition velocities to zero. Plates A, B and C show runs with NO emissions of 5.6×10^9 , 5.6×10^{10} and 5.6×10^{11} molecules $\text{cm}^{-2} \text{s}^{-1}$ respectively.

agreement for NO₂ is somewhat unexpected as the NO₂ deposition velocity is not insubstantial and NO₂ displays a clear concentration maximum close to the ground in the 1-D model, leading to an expectation of lower [NO₂] in the 1-D model. However NO₂ is mainly formed in the lowest 30 m, where substantial concentrations of NO exist. The K-profile (Fig. 2) shows that transport upwards, away from the deposition sink, is much more rapid than transport downwards towards the deposition sink, especially in the lowermost part of the NBL.

In contrast to the close inter-model agreement for NO_x species, the CiTTYCAT ozone mixing ratio is 57% smaller than the ozone mixing ratio in the 1-D model. The 1-D model surface ozone concentration is actually very close to that generated by CiTTYCAT, but the O₃ deposition flux in the 1-D model is reduced by the slow transport to the surface, leaving much higher O₃ concentrations at upper levels. Unlike NO₂, ozone does not have a night-time source in the NBL. The importance of deposition is

demonstrated in Fig. 7, by carrying out identical runs with each model with deposition turned off. Good agreement is then found between the 1-D model and CiTTYCAT for NO, NO_y and O₃. The apparent poor agreement for NO₂ is due to the formation of isoprene nitrates in CiTTYCAT and may be ignored for the purposes of this analysis (see Supplementary Information).

Sensitivity tests were carried out, varying the ozone deposition velocity in CiTTYCAT in order to attain the best possible agreement of CiTTYCAT and 1-D model O₃ concentrations. As there is only one independent variable, optimisation was carried out by eye until a clear best fit was found. This best fit was achieved using a reduced O₃ deposition velocity of 0.21 cm s⁻¹; a change of -65%. This change yields only a small reduction in NO₂ mixing ratio to 344 pptv, but suppresses the NO mixing ratio to 2.66 pptv. However, because of the short chemical lifetime of NO, the concentrations of NO carried over to the following day are not of great consequence.

The high NO concentrations seen near the ground in the 1-D model will modify night-time radical chemistry near the surface. A study of this is beyond the scope of this work. In a study of the vertical variability of night-time HO_x chemistry, Geyer and Stutz (2004b) found that elevated [NO] near the surface resulted in a substantial production of HO_x when VOCs were present. However, this scenario contained NO concentrations an order of magnitude larger than modelled here for OP3-1, and it is not expected that night-time radical chemistry caused by a near-surface build-up of NO will have an important effect on overall NBL oxidation for this low-[NO_x] scenario. Hence, whilst it may be concluded that it is not possible to change the set-up of the CiTTYCAT parameters to exactly replicate the 1-D model concentrations in the NBL, paying careful consideration to the night-time deposition velocity assignment in box models to take account of the variation in concentration across the NBL, may produce NO₂ and O₃ concentrations very similar to those in a 1-D model. We generalise this result in section 6.3, below, but first consider higher NO_x scenarios.

6.2 Medium to high [NO_x] conditions

The comparison between CiTTYCAT and the 1-D model is repeated with the same model set up, but with NO emissions increased by an order of magnitude (Fig. 7). As such an increase NO emission would result in a greater ozone production potential, the initial O₃ mixing ratio is set at an increased value of 60 ppbv, in line with studies measuring rural ozone concentrations close to an urban NO_x source (e.g. Murphy et al, 2007; Fuentes et al, 2007). The inter-model comparison of mixing ratios of NO₂, NO, NO_y and O₃ in Fig. 7 shows a similar level of agreement between the box and 1-D model as was seen in the runs under low-[NO_x] conditions. Only O₃ mixing ratios show a substantial discrepancy between the models, although the NO discrepancy has also increased slightly. Once again, removing dry deposition as a sink, largely eliminates the difference in the end-of-night mixing ratios for NO_y and O₃ between the box and 1-D models.

As for the low-[NO_x] scenario, reducing the O₃ deposition velocity in CiTTYCAT, this time to 0.19 cm s⁻¹, gives a very good fit to the 1-D model O₃ concentration. Compared to the low-[NO_x] scenario, the departure for NO₂ is increased, with CiTTYCAT NO₂ mixing ratios now 24% less than those in the 1-D model. Furthermore, in this mid-[NO_x] scenario, NO₃ mixing ratios varying from <1 pptv at 1 m to >200 pptv at 200 m are generated in the 1-D model. When combined with a likely layer of high [OH] near the surface, following Geyer and Stutz (2004b), there is clearly scope for considerable departures in oxidation behaviour between CiTTYCAT and the 1-D model.

Therefore, whilst the approach of modifying deposition velocities in the box model may be sufficient to effectively model daybreak NBL concentrations of O₃, NO and NO₂ in this scenario, box modelled daybreak concentrations of other species may only be treated with confidence at low [NO_x], such as that observed during OP3-1.

Increasing NO emissions by an additional order of magnitude to 5.6×10^{11} molecules cm⁻² s⁻¹, yields very substantial differences between the two models, even when dry deposition is turned off. These differences are particularly evident in the ozone concentration; the high [NO] in this scenario allows rapid ozone loss to proceed via reaction R1. However in the 1-D model NO is concentrated near the surface, leading to near-total depletion of O₃ near the surface, but the rate of ozone loss in the 1-D model is limited by the rate of transport of ozone to the surface, leading to substantial differences in ozone loss rate, and hence ambient [O₃], compared to the box model.

6.3 Generalised fit for low [NO_x] conditions

The box model is clearly inappropriate for modelling NO_x and O₃ chemistry in the stable NBL under high [NO_x] conditions. The only way to account for such modifications of the chemistry would be to define effective reaction rates, following e.g. Galmarini et al (1997). However such parameterisations require detailed knowledge of the distribution of species within the boundary layer, making this approach tied to either 1-D modelling or vertically-resolved measurements throughout the NBL. In section 6.1 it was shown that, under low [NO_x], it is possible to impose agreement between the box model and the 1-D model by modifying the ozone deposition velocity in the box model. This approach extends the deposition velocity definition to include the reduced mixing within the NBL, resulting in an effective deposition velocity. As the modification to the deposition velocity required will clearly depend upon the *K* profile used, an empirical relation is developed to describe this dependence. The theory used to describe the *K* profile in section 5 (Seinfeld and Pandis, 2006) produces a characteristic profile shape for a stable atmosphere, therefore what is of interest here is the magnitude of the *K* profile. From equations 5 and 6 the *K* profile is proportional to $u^* \times L$. As u^* and L are commonly calculated statistics to describe the atmospheric stability, it is convenient to use them to describe the effective mixing rate in the box model.

Figure 8 shows how the fitted ozone deposition velocity for the box model (V_b) varies with the natural log of $u^* \times L$, when the ozone deposition velocity in the 1-D model (V_c) is 0.6 cm s⁻¹. The equation of the line in Fig 8.,

$$V_b = 0.095 \ln(u^* \times L) + 0.27, \quad (7)$$

only holds when $V_c = 0.6$ cm s⁻¹. However, by repeating the process with V_c varying from 0.4 - 1.2 cm s⁻¹, expressions for both the gradient and the intercept were developed by plotting them against V_c . We can therefore relate the modified (V_b) and unmodified (V_c) deposition velocity:

$$V_b = m \ln(u^* \times L) + c \quad 0.4 \leq V_c \leq 1.2 \text{ cm s}^{-1}, [\text{NO}_x] < 1 \text{ ppbv}, [\text{O}_3] > 1 \text{ ppbv} \quad (8)$$

Where the gradient, m , is

$$m = 0.066 \ln(V_c) + 0.13 \quad (9)$$

and the intercept, c , is,

$$c = 0.12 \ln(V_c) + 0.33 \quad (10)$$

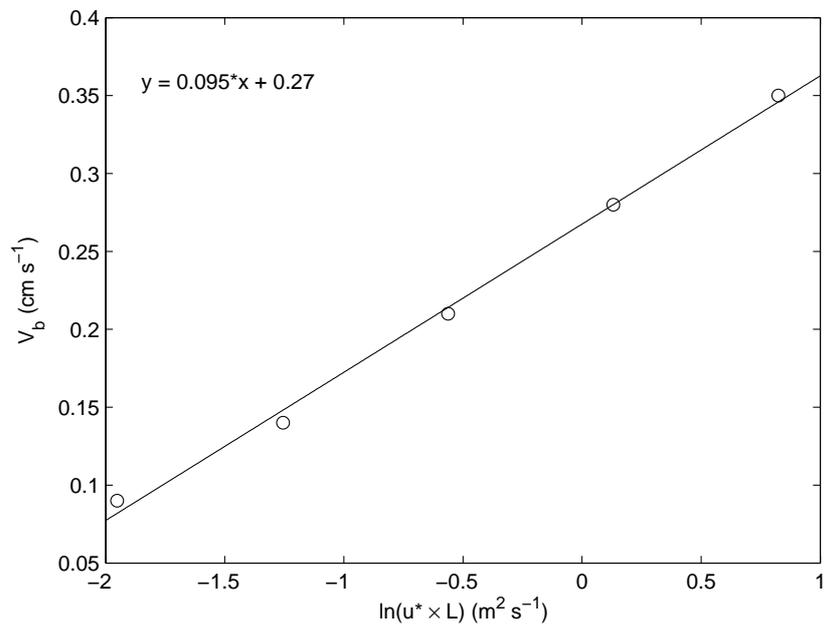


Figure 8: Modified ozone deposition velocity for the box model (V_b) as a function of the natural log of $u^* \times L$, when the ozone deposition velocity in the 1-D model (V_c) is 0.6 cm s^{-1} .

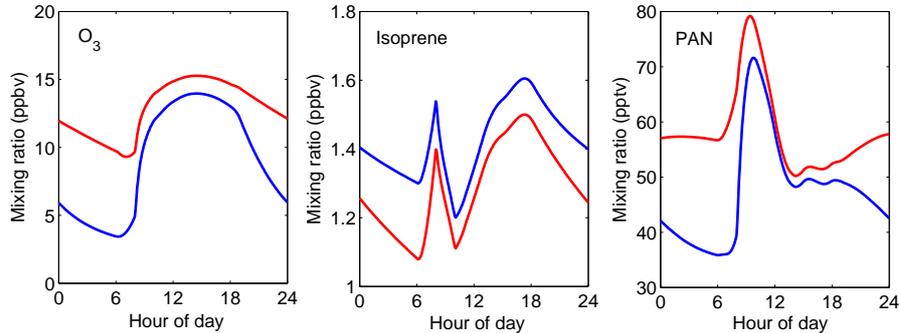


Figure 9: Result of using an effective O_3 deposition velocity (65% reduction) in CiTTyCAT (red line), compared to the default CiTTyCAT run shown in Fig. 1 (blue line) for O_3 (left), isoprene (centre) and PAN (right).

The friction velocity, u^* , is given in units of ms^{-1} , and the Monin-Obukhov length, L , is in units of m. Equations 8 to 10 can be used to estimate an appropriate adjustment for the night-time deposition of ozone in box models in order to account for the effects of limited mixing, under low $[\text{NO}_x]$ conditions.

It is important to note that the low-to-medium $[\text{NO}_x]$ conditions under which equation 8 holds are influenced by the emissions, the deposition rate and the mixing height. Therefore it is the ambient concentrations which define low NO_x , not the emission rate. Low $[\text{NO}_x]$ conditions are considered to be those where the NO_x mixing ratio is <1 ppbv. Additionally equation 8 does not hold in situations where O_3 is almost completely depleted at night, either by deposition or by NO titration.

Applying equations 8-10 for the conditions prevalent during OP3 yields a V_b which is 65% smaller than V_c . To test the importance of using an effective deposition velocity, the CiTTyCAT model run shown in Fig. 1 was repeated using this value of V_b . The result is compared with the original model run in Fig. 9. No comparison with measurements is made here as the original model run was optimised to agree with the daytime measurements by varying the emissions (see Pugh et al, 2010), therefore a measurement comparison would be misleading in this instance. A 9% increase in the peak daytime ozone concentration is calculated. The change in the oxidant budget leads to a 7% reduction in the peak isoprene mixing ratio and an 11% increase in the peak PAN mixing ratio. The sharp drop in PAN at midday is due to the loss of reactive nitrogen (particularly isoprene nitrates) via wet deposition, reducing PAN formation. Although, as mentioned previously, night-time rainfall was limited during the study period, afternoon rainfall was often substantial (see Hewitt et al, 2010). On days where wet deposition does not occur the increase in PAN is likely to be larger.

6.4 Application of O_3 parametrisation in other models

The O_3 deposition velocity parametrisation developed in Section 6.3 may be directly applied in other box models studying the boundary layer under low- NO_x conditions. The parametrisation could also be applied to global climate-chemistry models with limited vertical resolution near the surface, if they use a bulk mixing approach within the NBL. For those global climate-chemistry models which use a more complex ap-

proach to modelling the NBL, between the complexity of the box model and 1-D model approaches used in this paper (for instance a K profile with a limited number of levels), the conclusions of this work will still apply. In these cases we suggest that modellers must use their own judgement to assess whether their PBL scheme is capable of capturing sufficiently the type of concentration gradients resolved by the 1-D model herein. Should the model resolution be too coarse to capture these gradients, we suggest applying the parametrisation only in the lowermost layer of the model, where the concentration gradient of ozone is steepest. The parameters L and u^* must be available within the model. We recommend that any implementation in a global model is tested against the case presented in this paper, using a single-column implementation of the model to cover the NBL. We have included in the Supplementary Information details of the 1-D model concentrations with which satisfactory agreement must be achieved.

7 Summary

The CiTTyCAT box model is unable to replicate night-time measurements made at 5 m above ground level due to stratification within the nocturnal boundary layer. However a one-dimensional vertical transport model is able to reproduce many of the salient features of the measurements, especially the elevated 5-m NO concentrations in conjunction with relatively low NO₂ concentrations (compared to the [NO₂] expected based upon the NO and O₃ concentrations alone).

Comparison of average NBL concentrations of NO, NO₂ and O₃ between CiTTyCAT and the 1-D model show that CiTTyCAT is able to generate reasonable values for these tracers at the end of the night under low [NO_x] conditions. However the fit to [O₃] can be greatly improved by taking account for the effect of the nocturnal vertical concentration profile of O₃ on its deposition velocity. A generalised equation to define an effective night-time O₃ deposition velocity for box models is presented. Use of an effective O₃ deposition velocity during the night results in notable changes in concentrations during the following day for important species for atmospheric chemistry, such as ozone, isoprene and PAN. However, under high [NO_x], the vertical variation in chemistry renders the box model inappropriate for modelling night-time concentrations.

The parameterisation we have developed in section 6.3 would also be suitable for global climate-chemistry models with limited vertical resolution near the surface, if they use a bulk mixing approach within the NBL. For those global climate-chemistry models which use a more complex approach to modelling the NBL, between the complexity of the box model and 1-D model approaches used in this paper (for instance a K profile with a limited number of levels), the conclusions of this work will still apply. However a different correcting parameterisation for the ozone deposition velocity at low [NO_x] would need to be formulated according to the specific characteristics of the individual model.

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References

- Acevedo OC, Moraes OLL, Degrazia GA, Medeiros LE (2006) Intermittency and the exchange of scalars in the nocturnal surface layer. *Bound-Lay Meteorol* 119(1):41–55, DOI 10.1007/s10546-005-9019-3
- ApSimon H, Barker B, Kayin S (1994) Modelling studies of the atmospheric release and transport of ammonia in anticyclonic episodes. *Atmos Environ* 28(4):665–678
- Aumont B, Chervier F, Laval S (2003) Contribution of HONO sources to the $\text{NO}_x/\text{HO}_x/\text{O}_3$ chemistry in the polluted boundary layer. *Atmos Environ* 37(4):487–498
- Biesenthal T, Bottenheim J, Shepson P, Brickell P (1998) The chemistry of biogenic hydrocarbons at a rural site in eastern Canada. *J Geophys Res - Atmos* 103(D19):25,487–25,498, International Workshop on Biogenic Hydrocarbons in the Atmospheric Boundary Layer, Charlottesville, Virginia, Aug, 1997
- Brost R, Delany A, Huebert B (1988) Numerical modeling of concentrations and fluxes of HNO_3 , NH_3 , and NH_4NO_3 near the surface. *J Geophys Res - Atmos* 93(D6):7137–7152
- Brown SS, Dube WP, Osthoff HD, Wolfe DE, Angevine WM, Ravishankara AR (2007) High resolution vertical distributions of NO_3 and N_2O_5 through the nocturnal boundary layer. *Atmos Chem Phys* 7:139–149
- Businger J, Wyngaard J, Izumi Y, Bradley E (1971) Flux-Profile Relationships in Atmospheric Surface Layer. *J Atmos Sci* 28(2):181–&
- Coyle M, Fowler D, Nemitz E, Philips G, Storeton-West R, Thomas R (2006) Field measurements of the ozone flux to vegetation. In: *Ozone Umbrella: Effects of Ground-level Ozone on (Upland) Vegetation in the UK.*, Centre of Ecology and Hydrology, UK, CEH C02158 Report No. AS 06/02., p 68104
- Cros B, Fontan J, Minga A, Helas G, Nganga D, Delmas R, Chapuis A, Benech B, Druilhet A, Andreae M (1992) Vertical profiles of ozone between 0 metres and 400 metres in and above the African equatorial forest. *J Geophys Res-Atm* 97(D12):12,877–12,887
- Daescu D, Sandu A, Carmichael G (2003) Direct and adjoint sensitivity analysis of chemical kinetic systems with kpp: Ii – validation and numerical experiments. *Atmospheric Environment* 37:5097–5114

- Damian V, Sandu A, Damian M, Potra F, Carmichael G (2002) The kinetic preprocessor kpp – a software environment for solving chemical kinetics. *Computers and Chemical Engineering* 26(11):1567–1579
- Donovan R, Hope E, Owen S, Mackenzie A, Hewitt C (2005) Development and application of an urban tree air quality score for photochemical pollution episodes using the birmingham, united kingdom, area as a case study. *Environ Sci Technol* 39:6730–6738
- Emmerson K, MacKenzie A, Owen S, Evans M, Shallcross D (2004) A lagrangian model with simple primary and secondary aerosol scheme 1: comparison with uk pm₁₀ data. *Atmos Chem Phys* 4:2161–2170
- Evans M, Shallcross D, Law K, Wild J, Simmonds P, Spain T, Berrisford P, Methven J, Lewis A, McQuaid J, Pilling M, Bandy B, Penkett S, Pyle J (2000) Evaluation of a Lagrangian box model using field measurements from EASE (Eastern Atlantic Summer Experiment) 1996. *Atmos Environ* 34(23):3843–3863
- Faloona I, Tan D, Brune W, Hurst J, Barket D, Couch T, Shepson P, Apel E, Riemer D, Thornberry T, Carroll M, Sillman S, Keeler G, Sagady J, Hooper D, Paterson K (2001) Nighttime observations of anomalously high levels of hydroxyl radicals above a deciduous forest canopy. *J Geophys Res - Atmos* 106(D20):24,315–24,333
- Fan S, Wofsy S, Bakwin P, Jacob D, Fitzjarrald D (1990) Atmosphere-biosphere exchange of CO₂ and O₃ in the Central-Amazon-Forest. *J Geophys Res - Atmos* 95(D10):16,851–16,864
- Fitzjarrald D, Lenschow D (1983) Mean concentration and flux profiles for chemically reactive species in the atmospheric surface-layer. *Atmos Environ* 17(12):2505–2512
- Fitzjarrald D, Moore K (1990) Mechanisms of nocturnal exchange between the rain-forest and the atmosphere. *J Geophys Res - Atmos* 95(D10):16,839–16,850
- Fowler D, Pilegaard K, Sutton MA, Ambus P, Raivonen M, Duyzer J, Simpson D, Fagerli H, Fuzzi S, Schjoerring JK, Granier C, Neftel A, Isaksen ISA, Laj P, Maione M, Monks PS, Burkhardt J, Daemmgen U, Neirynek J, Personne E, Wichink-Kruit R, Butterbach-Bahl K, Flechard C, Tuovinen JP, Coyle M, Gerosa G, Loubet B, Altimir N, Gruenhage L, Ammann C, Cieslik S, Paoletti E, Mikkelsen TN, Rø-Poulsen H, Cellier P, Cape JN, Horvath L, Loreto F, Niinemets U, Palmer PI, Rinne J, Misztal P, Nemitz E, Nilsson D, Pryor S, Gallagher MW, Vesala T, Skiba U, Brüeggemann N, Zechmeister-Boltenstern S, Williams J, O’Dowd C, Facchini MC, de Leeuw G, Flossman A, Chaumerliac N, Erismann JW (2009) Atmospheric composition change: Ecosystems-Atmosphere interactions. *Atmos Environ* 43(33, Sp. Iss. SI):5193–5267, DOI 10.1016/j.atmosenv.2009.07.068
- Fuentes JD, Wang D, Bowling DR, Potosnak M, Monson RK, Goliff WS, Stockwell WR (2007) Biogenic hydrocarbon chemistry within and above a mixed deciduous forest. *J Atmos Chem* 56(2):165–185, DOI 10.1007/s10874-006-9048-4
- Galmarini S, Duynkerke P, deArellano I (1997) Evolution of nitrogen oxide chemistry in the nocturnal boundary layer. *J Appl Meteorol* 36(7):943–957

- Gao W, Wesely M (1994) Numerical modeling of the turbulent fluxes of chemically reactive trace gases in the atmospheric boundary-layer. *J Appl Meteorol* 33(7):835–847
- Gao W, Wesely M, Doskey P (1993) Numerical modeling of the turbulent-diffusion and chemistry of NO_x, O₃, isoprene, and other reactive trace gases in and above a forest canopy. *J Geophys Res-Atmos* 98(D10):18,339–18,353
- Geyer A, Stutz J (2004a) Vertical profiles of NO₃, N₂O₅, O₃, and NO_x in the nocturnal boundary layer: 2. Model studies on the altitude dependence of composition and chemistry. *J Geophys Res - Atmos* 109(D12), DOI 10.1029/2003JD004211
- Geyer A, Stutz J (2004b) The vertical structure of OH-HO₂-RO₂ chemistry in the nocturnal boundary layer: A one-dimensional model study. *J Geophys Res* 109(D16), DOI 10.1029/2003JD004425
- Glaser K, Vogt U, Baumbach G, Volz-Thomas A, Geiss H (2003) Vertical profiles of O₃, NO₂, NO_x, VOC, and meteorological parameters during the Berlin Ozone Experiment (BERLIOZ) campaign. *J Geophys Res-Atm* 108(D4), DOI 10.1029/2001JD002475
- Guenther A, Karl T, Harley P, Wiedinmyer C, Palmer PI, Geron C (2006) Estimates of global terrestrial isoprene emissions using MEGAN (Model of Emissions of Gases and Aerosols from Nature). *Atmos Chem Phys* 6:3181–3210
- Gusten H, Heinrich G, Sprung D (1998) Nocturnal depletion of ozone in the Upper Rhine Valley. *Atmos Environ* 32(7):1195–1202
- Hamba F (1987) Statistical-analysis of chemically reacting passive scalars in turbulent shear flows. *J Phys Soc Jpn* 56(1):79–96
- Hamer P, Shallcross D (2007) Modelling the impact of oxygenated VOC and meteorology upon the boundary layer photochemistry at the South Pole. *Atmos Sci Lett* 8:14–20
- Hewitt CN, Lee JD, MacKenzie AR, Barkley MP, Carslaw N, Carver GD, Chappell NA, Coe H, Collier C, Commane R, Davies F, Davison B, DiCarlo P, Di Marco CF, Dorsey JR, Edwards PM, Evans MJ, Fowler D, Furneaux KL, Gallagher M, Guenther A, Heard DE, Helfter C, Hopkins J, Ingham T, Irwin M, Jones C, Karunaharan A, Langford B, Lewis AC, Lim SF, MacDonald SM, Mahajan AS, Malpass S, McFiggans G, Mills G, Misztal P, Moller S, Monks PS, Nemitz E, Nicolas-Perea V, Oetjen H, Oram DE, Palmer PI, Phillips GJ, Pike R, Plane JMC, Pugh T, Pyle JA, Reeves CE, Robinson NH, Stewart D, Stone D, Whalley LK, Yin X (2010) Overview: oxidant and particle photochemical processes above a south-east asian tropical rainforest (the op3 project): introduction, rationale, location characteristics and tools. *Atmos Chem Phys* 10(1):169–199, URL <http://www.atmos-chem-phys.net/10/169/2010/>
- Hewitt CN, MacKenzie AR, Di Carlo P, Di Marco CF, Dorsey JR, Evans M, Fowler D, Gallagher MW, Hopkins JR, Jones CE, Langford B, Lee JD, Lewis AC, Lim SF, McQuaid J, Misztal P, Moller SJ, Monks PS, Nemitz E, Oram DE, Owen SM, Phillips GJ, Pugh TAM, Pyle JA, Reeves CE, Ryder J, Siong J, Skiba U, Stewart DJ (2009) Nitrogen management is essential to prevent tropical oil palm plantations

- from causing ground-level ozone pollution. *Proc Nat Acad Sci USA* 106(44):18,447–18,451, DOI 10.1073/pnas.0907541106
- Horowitz LW, Fiore AM, Milly GP, Cohen RC, Perring A, Wooldridge PJ, Hess PG, Emmons LK, Lamarque JF (2007) Observational constraints on the chemistry of isoprene nitrates over the eastern United States. *J Geophys Res* 112(D12), DOI 10.1029/2006JD007747
- Hov O (1983) One-dimensional vertical model for ozone and other gases in the atmospheric boundary-layer. *Atmos Environ* 17(3):535–549
- Kerkweg A, Buchholz J, Ganzeveld L, Pozzer A, Tost H, Joeckel P (2006) Technical note: An implementation of the dry removal processes DRY DEPosition and SED-Imentation in the modular earth submodel system (MESSy). *Atmos Chem Phys* 6:4617–4632
- Kleffmann J (2007) Daytime sources of nitrous acid (HONO) in the atmospheric boundary layer. *ChemPhysChem* 8(8):1137–1144, DOI 10.1002/cphc.200700016
- Kleffmann J, Becker K, Wiesen P (1998) Heterogeneous NO₂ conversion processes on acid surfaces: Possible atmospheric implications. *Atmos Environ* 32(16):2721–2729
- Lowe D, Topping D, McFiggans G (2009) Modelling multi-phase halogen chemistry in the remote marine boundary layer: investigation of the influence of aerosol size resolution on predicted gas- and condensed-phase chemistry. *Atmos Chem Phys* 9(14):4559–4573
- Matsuda K, Watanabe I, Wingpud V (2005) Ozone dry deposition above a tropical forest in the dry season in northern Thailand. *Atmos Environ* 39(14):2571–2577, DOI 10.1016/j.atmosenv.2005.01.011
- Michou M, Laville P, Serca D, Fotiadi A, Bouchou P, Peuch V (2005) Measured and modeled dry deposition velocities over the ESCOMPTE area. *Atmos Res* 74(1-4):89–116, DOI 10.1016/j.atmosres.2004.04.011
- Mihele C, Hastie D (2003) Radical chemistry at a forested continental site: Results from the PROPHET 1997 campaign. *J Geophys Res - Atmos* 108(D15), DOI 10.1029/2002JD002888
- Monks P (2005) Gas-phase radical chemistry in the troposphere. *Chemical Society Reviews* 34(5):376–395, DOI 10.1039/b307982c
- Murphy JG, Day DA, Cleary PA, Wooldridge PJ, Millet DB, Goldstein AH, Cohen RC (2007) The weekend effect within and downwind of Sacramento - Part 1: Observations of ozone, nitrogen oxides, and VOC reactivity. *Atmos Chem Phys* 7(20):5327–5339
- Pearson G, Davies F, Collier C (2010) Remote sensing of the tropical rain forest boundary layer using pulsed doppler lidar. *Atmos Chem Phys Discuss* 10(2):5021–5049, URL <http://www.atmos-chem-phys-discuss.net/10/5021/2010/>
- Pike R, Lee J, Young P, Moller S, Carver G, Yang X, Misztal P, Langford B, Stewart D, Reeves C, Hewitt C, Pyle J (2009) Can a gcm mechanism reproduce no, no₂ and o₃ data above a tropical rainforest? *Atmos Chem Phys Discuss* pp 27,611–27,648

- Pisano J, McKendry I, Steyn D, Hastie D (1997) Vertical nitrogen dioxide and ozone concentrations measured from a tethered balloon in the Lower Fraser Valley. *Atmos Environ* 31(14):2071–2078
- Pugh TAM, MacKenzie AR, Hewitt CN, Langford B, Edwards PM, Furneaux KL, Heard DE, Hopkins JR, Jones CE, Karunaharan A, Lee J, Mills G, Misztal P, Moller S, Monks PS, Whalley LK (2010) Simulating atmospheric composition over a South-East Asian tropical rainforest: performance of a chemistry box model. *Atmos Chem Phys* 10(1):279–298
- Read KA, Mahajan AS, Carpenter LJ, Evans MJ, Faria BVE, Heard DE, Hopkins JR, Lee JD, Moller SJ, Lewis AC, Mendes L, McQuaid JB, Oetjen H, Saiz-Lopez A, Pilling MJ, Plane JMC (2008) Extensive halogen-mediated ozone destruction over the tropical Atlantic Ocean. *Nature* 453(7199):1232–1235, DOI 10.1038/nature07035
- Real E, Law KS, Weinzierl B, Fiebig M, Petzold A, Wild O, Methven J, Arnold S, Stohl A, Huntrieser H, Roiger A, Schlager H, Stewart D, Avery M, Sachse G, Browell E, Ferrare R, Blake D (2007) Processes influencing ozone levels in Alaskan forest fire plumes during long-range transport over the North Atlantic. *J Geophys Res - Atmos* 112(D10), DOI 10.1029/2006JD007576
- Real E, Law KS, Schlager H, Roiger A, Huntrieser H, Methven J, Cain M, Holloway J, Neuman JA, Ryerson T, Flocke F, de Gouw J, Atlas E, Donnelly S, Parrish D (2008) Lagrangian analysis of low altitude anthropogenic plume processing across the North Atlantic. *Atmos Chem Phys* 8(24):7737–7754
- Rummel U, Ammann C, Kirkman GA, Moura MAL, Foken T, Andreae MO, Meixner FX (2007) Seasonal variation of ozone deposition to a tropical rain forest in south-west Amazonia. *Atmos Chem Phys* 7(20):5415–5435
- Ryder J, Lowe D, McFiggans G, Nemitz E (2010) Application of a coupled chemistry and deposition model to simulate the effect of gas / aerosol partitioning on size-segregated aerosol fluxes above and within a heathland canopy. in preparation
- Sandu A, Daescu D, Carmichael G (2003) Direct and adjoint sensitivity analysis of chemical kinetic systems with kpp: I – theory and software tools. *Atmospheric Environment* 37:5083–5096
- Seinfeld J, Pandis S (2006) *Atmospheric chemistry and physics: From air pollution to climate change*, 2nd edn. Wiley-interscience, New Jersey, USA
- Strong J, Whyatt JD, Hewitt CN, Derwent RG (2009) Development and application of a Lagrangian model to determine the origins of ozone episodes in the UK. *Atmos Environ*, in press DOI 10.1016/j.atmosenv.2009.11.019
- Stutz J, Alicke B, Ackermann R, Geyer A, White A, Williams E (2004) Vertical profiles of NO₃, N₂O₅, O₃, and NO_x in the nocturnal boundary layer: 1. Observations during the Texas Air Quality Study 2000. *J Geophys Res-Atmos* 109(D16), DOI 10.1029/2004JD005216
- Thompson A, Lenschow D (1984) Mean profiles of trace reactive species in the unpolluted marine surface-layer. *J Geophys Res - Atmos* 89(ND3):4788–4796

- Toyota K, Kanaya Y, Takahashi M, Akimoto H (2004) A box model study on photochemical interactions between VOCs and reactive halogen species in the marine boundary layer. *Atmos Chem Phys* 4:1961–1987
- Trebs I, Lara L, Zeri L, Gatti L, Artaxo P, Dlugi R, Slanina J, Andreae M, Meixner F (2006) Dry and wet deposition of inorganic nitrogen compounds to a tropical pasture site (Rondonia, Brazil). *Atmos Chem Phys* 6:447–469
- Vila-Guerau de Arellano J (2003) Bridging the gap between atmospheric physics and chemistry in studies of small-scale turbulence. *B Am Meteorol Soc* 84(1):51–56, DOI 10.1175/BAMS-84-1-51
- Vinuesa J, Vilà-Guerau de Arellano J (2005) Introducing effective reaction rates to account for the inefficient mixing of the convective boundary layer. *Atmos Environ* 39(3):445–461, DOI 10.1016/j.atmosenv.2004.10.003
- Watt S, Wagner-Riddle C, Edwards G, Vet R (2004) Evaluating a flux-gradient approach for flux and deposition velocity of nitrogen dioxide over short-grass surfaces. *Atmos Environ* 38(17):2619–2626, DOI 10.1016/j.atmosenv.2004.02.021
- Wayne R, Barnes I, Biggs P, Burrows J, Canosamas C, Hjorth J, Lebras G, Moortgat G, Perner D, Poulet G, Restelli G, Sidebottom H (1991) The nitrate radical - physics, chemistry, and the atmosphere. *Atmos Environ A-Gen* 25(1):1–203
- Wesely M, Hicks B (2000) A review of the current status of knowledge on dry deposition. *Atmos Environ* 34(12-14):2261–2282
- Wild O, Law K, McKenna D, Bandy B, Penkett S, Pyle J (1996) Photochemical trajectory modeling studies of the North Atlantic region during August 1993. *J Geophys Res* 101(D22):29,269–29,288
- Yienger J, Levy H (1995) Empirical model of global soil biogenic NO_x emissions. *J Geophys Res* 100(D6):11,447–11,464
- Zhang L, Vet R, O'Brien JM, Mihele C, Liang Z, Wiebe A (2009) Dry deposition of individual nitrogen species at eight Canadian rural sites. *J Geophys Res - Atmos* 114, DOI 10.1029/2008JD010640