

Sintering and oxidation of GdB₄ synthesized by B₄C reduction method

Sonber, Jitendra Kumar; Murthy, Tammana Shri Ram Chandra; Sairam, Kannan; Kain, Vivekanand

DOI:

[10.4191/kcers.2017.54.2.02](https://doi.org/10.4191/kcers.2017.54.2.02)

License:

Creative Commons: Attribution (CC BY)

Document Version

Publisher's PDF, also known as Version of record

Citation for published version (Harvard):

Sonber, JK, Murthy, TSRC, Sairam, K & Kain, V 2017, 'Sintering and oxidation of GdB₄ synthesized by B₄C reduction method', *Journal of the Korean Ceramic Society*, vol. 54, no. 2, pp. 121-127. ⁴
<https://doi.org/10.4191/kcers.2017.54.2.02>

[Link to publication on Research at Birmingham portal](#)

Publisher Rights Statement:

Checked for eligibility: 28/06/2018

General rights

Unless a licence is specified above, all rights (including copyright and moral rights) in this document are retained by the authors and/or the copyright holders. The express permission of the copyright holder must be obtained for any use of this material other than for purposes permitted by law.

- Users may freely distribute the URL that is used to identify this publication.
- Users may download and/or print one copy of the publication from the University of Birmingham research portal for the purpose of private study or non-commercial research.
- User may use extracts from the document in line with the concept of 'fair dealing' under the Copyright, Designs and Patents Act 1988 (?)
- Users may not further distribute the material nor use it for the purposes of commercial gain.

Where a licence is displayed above, please note the terms and conditions of the licence govern your use of this document.

When citing, please reference the published version.

Take down policy

While the University of Birmingham exercises care and attention in making items available there are rare occasions when an item has been uploaded in error or has been deemed to be commercially or otherwise sensitive.

If you believe that this is the case for this document, please contact UBIRA@lists.bham.ac.uk providing details and we will remove access to the work immediately and investigate.

Sintering and Oxidation of GdB_4 Synthesized by B_4C Reduction Method

Jitendra Kumar Sonber[†], Tammana Shri Ram Chandra Murthy,
Kannan Sairam, and Vivekanand Kain

Materials Processing and Corrosion Engineering Division, Bhabha Atomic Research Centre, Mumbai 400085, India

(Received October 21, 2016; Revised January 18, January 19, 2017; Accepted January 19, 2017)

ABSTRACT

Gadolinium tetraboride (GdB_4) was synthesized by reduction of Gd_2O_3 using boron carbide in presence of carbon. Effect of temperature on product quality was investigated. Pure GdB_4 powder was obtained in vacuum at $1500^\circ C$. Pressureless sintering experiments revealed that sintering takes place only above $1600^\circ C$. A maximum density of 77.1% of the theoretical value was obtained at $1800^\circ C$ by pressureless sintering. Hot pressing resulted in 95.5% of theoretical density at the lower temperature of $1700^\circ C$ under 35 MPa pressure. Hardness and fracture toughness of dense GdB_4 were measured and found to be 21.4 GPa and $2.3 \text{ MPa}\cdot\text{m}^{1/2}$, respectively. After exposure to air at $900^\circ C$, the formation of a porous and non-protective oxide layer was observed.

Key words : GdB_4 , Synthesis, Hot pressing, Boride, Oxidation

1. Introduction

Rare earth borides are exceptional ceramics characterized by high melting point, high hardness, excellent thermal stability, low vapour pressure, high electrical and thermal conductivity and low thermal expansion coefficient.¹⁾ Due to high neutron absorption cross section of boron, rare earth borides are candidate materials for neutron absorber applications in nuclear reactors.²⁻⁴⁾ They are good thermionic electron emitters due to their low work function, low vapour pressure, high brightness and long service life.⁵⁻⁷⁾ Though borides of Lanthanum (La) and Cerium (Ce) have been studied extensively, literature on Gadolinium boride has been minimal. The high neutron absorbing capacity of Gd and boron makes the resulting compound suitable for neutron absorber applications in nuclear reactors. In the Gd-B system, the GdB_4 phase is the most stable. Important properties of GdB_4 are listed in Table 1.

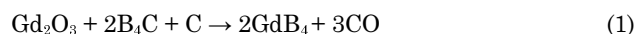
Most of the literature on GdB_4 has been concerned with its physical properties.⁸⁻¹⁰⁾ Reports on synthesis and densification of GdB_4 have been very limited.^{5,11,12)} Lazorenko *et al.*¹¹⁾ prepared GdB_4 by borothermic reduction of Gd_2O_3 at around $2000^\circ C$ in vacuum. Itoh *et al.*¹²⁾ prepared GdB_4 and GdB_6 by decomposition of the boron cage compound $Gd_2(B_{10}H_{10})_3$.

Gadolinium tetraboride powder is not easily available. In the present study, a process has been developed to prepare GdB_4 powder. In this process, GdB_4 was synthesized by boron carbide reduction of Gd_2O_3 in the presence of carbon, as per Eq. (1). We have not come across any report on the synthesis of GdB_4 powder by boron carbide reduction method.

Table 1. Properties of Gadolinium Tetraboride¹⁾

Property	GdB_4
Crystal structure	Tetragonal
Density (gm/cc)	6.47
Melting point ($^\circ C$)	2650
Thermal expansion coefficient ($^\circ C^{-1}$)	7.0×10^{-6}
Thermal conductivity ($WM^{-1}K^{-1}$)	148.5
Hardness (GPa)	18.6

This method is expected to be economical since the raw materials are Gd_2O_3 , B_4C and carbon, which are comparatively less expensive than elemental Boron and boron cage compounds used in previous studies.^{11,12)}



In many applications, borides are required to be in the form of dense shapes. Sintering of borides is a challenging task because the high melting point and low self diffusivity of these materials make sintering difficult.¹³⁻¹⁵⁾ Moreover, boride powders are generally covered with surface oxide layers that decrease the surface energy of the particles.¹³⁻¹⁵⁾ In this study, pressureless sintering and hot pressing techniques were investigated for obtaining GdB_4 in the form of dense pellets.

Many borides are susceptible to oxidation, which limits their application in high temperature oxidizing environments. In this study, an oxidation test of GdB_4 was carried out to explore the possibility of using this material in high temperature oxidizing environments.

This study reports on the results of investigations of the synthesis, consolidation and oxidation behaviour of GdB_4 . To the best of authors' knowledge, there is no published

[†]Corresponding author : Jitendra Kumar Sonber
E-mail : jitendra@barc.gov.in
Tel : +91-22-2559-5112 Fax : +91-22-2559-5151

report on densification and oxidation of GdB_4 .

2. Experimental Procedure

2.1. Starting material

Gd_2O_3 (99.9% purity, supplied by M/s Advanced Technology Materials, Mumbai), B_4C (78.5% B, 19.5% C, < 1% O, 0.02% Fe, 0.02% Si; 5.34 μm particle median diameter; supplied by M/s Boron Carbide India) and petroleum coke (99.4% purity; 13.9 μm median particle diameter, supplied by M/s Assam Carbon, India) were used as starting materials. The starting materials were characterized by X-ray diffraction (XRD) analysis using $\text{Cu K}\alpha$ ($\lambda = 1.5404 \text{ \AA}$) radiation; this method has a detection limit of 4 vol%. Fig. 1 presents the XRD patterns of the starting powders, and confirms that there is no impurity phase in the starting powders, within the detection limit of XRD.

2.2. Synthesis

For the synthesis of GdB_4 , weighed quantities of gadolinium oxide, boron carbide and carbon in stoichiometric ratios were mixed thoroughly in a motorized mortar and pestle. The powder mixture was then pelletized under a pressure of about 200 MPa to obtain green pellets of 12 mm diameter. The pellets were then placed in a graphite crucible and heated in an induction furnace at a fixed temperature between 1200 to 1500°C for dwelling period of 2 h under a dynamic vacuum of 2×10^{-5} mbar. The temperature of the charge was measured using a two-colour pyrometer with an accuracy of $\pm 20^\circ\text{C}$. After the dwelling period, the furnace was allowed to cool to room temperature under vacuum; then, the partially sintered pellets were taken out, crushed and ground to a fine size using a high energy tungsten car-

bide lined cup grinding mill.

Major phases of powders were identified by XRD using $\text{Cu K}\alpha$ ($\lambda = 1.5404 \text{ \AA}$) radiation. Carbon and oxygen were measured by combustion method and by inert gas fusion (IGF) method, respectively. Particle morphology was examined by scanning electron microscope (SEM).

2.3. Densification

In the pressureless sintering (PS) experiments, GdB_4 powder was cold compacted under a pressure of about 280 MPa to form pellets of 12 mm diameter and 5 mm height with a green density of $\sim 62\%$ of the theoretical density (ρ_{th}). These compacts were sintered at a temperature between 1200 to 1800°C under a dynamic vacuum of 2×10^{-5} mbar in an induction furnace for holding periods of 2 h. In the hot pressing experiments, fine powders were hot pressed under a vacuum of 2×10^{-5} mbar in a high density graphite die (17 mm hole) by applying 35 MPa pressure at 1700°C for 2 h. The densities of the pellets were measured by the Archimedes principle. The fracture surfaces of the dense pellets was examined by SEM. The samples were broken by hammering to obtain a fracture surface.

Densified samples were polished to a mirror finish using diamond powder of various grades from 15 to 0.25 μm in an auto polisher (laboforce-3, Struers). Microhardness was measured at a load of 100 g and dwell time of 10 sec. Indentation fracture toughness (K_{IC}) data were evaluated by crack length measurement of the crack pattern formed around the Vickers indents (using 10 kg load), adopting the model formulation proposed by Anstis *et al.*¹⁶ $K_{IC} = 0.016(E/H)^{1/2}P/c^{3/2}$, where E is the Young's modulus, H the Vickers's hardness, P the applied indentation load, and c the half crack length.

2.4. Oxidation

Isothermal oxidation study of the dense GdB_4 was carried out at a temperature of 900°C for different time intervals. A hot pressed pellet of diameter 17 mm was cut into thin slices of 3 mm thickness. All the surfaces of the cut sample were ground with emery papers (1/0, 2/0, 3/0, 4/0) and finally polished with diamond paste to achieve a 1 μm finish. Oxidation tests were conducted in a resistance heated furnace. In order to avoid oxidation during heating, the sample was directly inserted into the hot zone after the furnace had reached the set temperature. Sample was placed in an alumina crucible loaded inside the furnace. The sample was oxidized in air for different time intervals (4, 8, 16, 32) at the set temperature of 900°C. The sample was carefully weighed to an accuracy of 0.01 mg before and after exposure, to determine the weight change during the oxidation process. The morphology and nature of the oxide layer were established by observing the surface in a field emission scanning electron microscope (FE-SEM) with energy dispersive spectroscopy (EDS). The phases present on the oxidized surface were analysed by XRD.

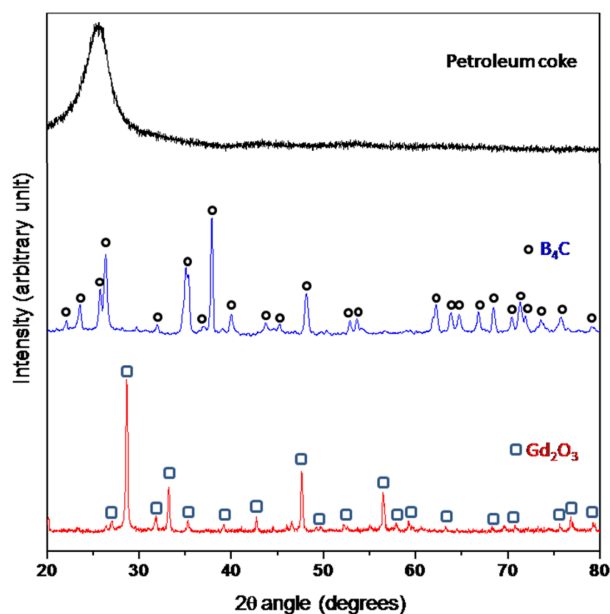
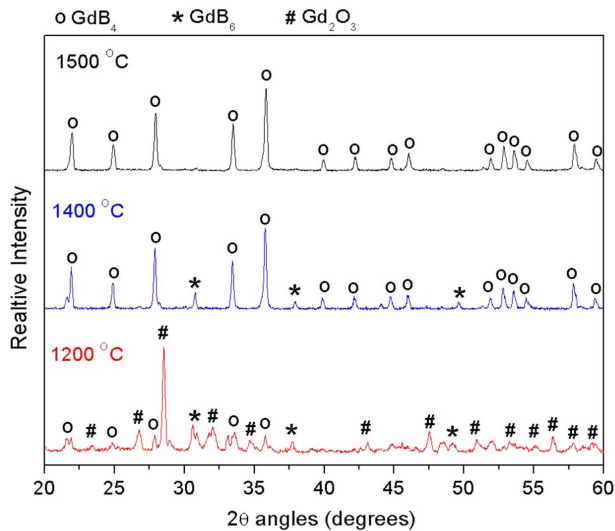


Fig. 1. XRD pattern of starting materials.

Table 2. Results of GdB₄ Synthesis Experiments

S. No.	Temp. (°C)	Weight loss (%)	Phases present
1	1200	5.7	Gd ₂ O ₃ , GdB ₄ , GdB ₆
2	1300	16.8	Gd ₂ O ₃ , GdB ₄ , GdB ₆
3	1400	18.2	GdB ₄ , GdB ₆
4	1500	18.6	GdB ₄

**Fig. 2.** XRD patterns of reaction product obtained at different temperatures.

3. Results and Discussion

3.1. Synthesis

Table 2 presents the processing parameters and results of synthesis experiments carried out in a vacuum induction furnace. At 1200 and 1300°C, weight loss noticed is about 5.7% and 16.8%, respectively, which values are lower than the theoretical weigh loss (17.3%), indicating that the reaction had started but did not complete. The product was analysed and found to contain Gd₂O₃, GdB₄ and GdB₆ phases. The presence of Gd₂O₃ confirms the incomplete reduction. Sample processed at 1400°C was observed to contain GdB₄

and GdB₆ phases. Only at a temperature of 1500°C was single phase GdB₄ obtained. The carbon and oxygen content of GdB₄ were measured and found to be 0.5 wt.% each.

Figure 2 presents the XRD patterns of the samples processed at different temperature. The XRD patterns reveal that the product obtained at 1500°C is pure GdB₄, whereas products obtained at lower temperatures also contain the Gd₂O₃ and GdB₆ phase. The GdB₆ phase could have formed via the following side reaction (reaction (2)).



At higher temperatures, the reaction kinetics for reaction (1) was fast and all the Gd₂O₃ was converted to GdB₄. In the Gd-B system, GdB₄ is the most stable phase. Free energy calculations for reactions (1) and (2) could not be done as thermodynamic data for GdB₄ and GdB₆ are not available.

Though the reaction starts at a lower temperature (< 1200°C), pure GdB₄ is obtained only at 1500°C. At low temperature, the reaction kinetics is very slow and hence a high temperature is required for faster diffusion of the solid reactant atoms through the product that formed at the interface.

The scientific literature on the processing of GdB₄ is scant, which was the motivation for this work. A summary of past studies is provided in Table 3.^{5,11,12)} The borothermic reduction process involves the use of elemental boron, which is expensive and pyrophoric. Preparation of GdB₄ by boron cage compounds and metal hydrides requires one additional step, as these materials are not readily available. Handling of metal hydride is also difficult due to its pyrophoric nature. Preparation by reaction between Gd₂O₃, B₄C and carbon seems to be the best process as it involves readily available charge material and does not have any handling issues. The process is also suitable for large scale production of boride powder. In this study, single phase GdB₄ was obtained.

Synthesis of other rare earth borides by this route has been reported in literature. Zaykoski *et al.* have prepared aYB₄ ceramic by this route at 1800°C in helium atmosphere.¹⁷⁾ Liu *et al.*¹⁸⁾ have reported that NdB₆ powder can be prepared by boron carbide reduction of Nd₂O₃ at 1500°C in vacuum. Xu *et al.*¹⁹⁾ have prepared LaB₆ by this route at 1650°C at atmospheric pressure. Sonber *et al.* have prepared LaB₆, EuB₆ and CeB₆ by this route.^{13,20,21)}

Table 3. Summary of Studies on GdB₄ Synthesis

S. No.	Reactants	Process Parameters	Product quality	Ref.
1	Gd ₂ O ₃ + B	Temp.: 1997°C Holding time: 1 h Vacuum: 10 ⁻² Pa (10 ⁻⁴ mbar)	Single phase GdB ₄	5
2	Gd ₂ (B ₁₀ H ₁₀) ₃ + GdH ₂	Temp: 1200-1400°C Atmosphere: Argon Time: 3 h	GdB ₄ and GdB ₆ phases	12
3	Gd ₂ O ₃ + B ₄ C	Temp.: 1470°C	GdB ₆ , GdB ₄	11
4	Gd ₂ O ₃ + B ₄ C + C	Temp.: 1500°C Vacuum: 10 ⁻⁵ mbar Time: 2 h	Single phase GdB ₄	Present Study

3.2. Densification and characterization

The particle size distribution of synthesized GdB_4 powder that was ground using a high energy vibratory cup mill is presented in Fig. 3, which shows a bimodal distribution of particles with peaks at 2.5 μm and 10 μm . Mean particle diameter was found to be 5.6 μm . An SEM image (Fig. 4) of the powder shows particle agglomerates of 4 - 7 μm size.

Results of pressureless sintering and hot pressing experiments are summarized in Table 4. Results of pressure less sintering at temperatures upto 1600°C have shown that the obtained density is close to the green density of 63%, and that no densification took place during sintering. At 1800°C, densification started and a density of 77.1% of the theoretical value was obtained. The presence of high porosity (~ 23 %) after sintering at 1800°C indicated that the densification of GdB_4 by pressureless sintering is extremely difficult. Hot pressing of GdB_4 at 1700°C under 35 MPa pressure resulted in a density that was 95.5% of the theoretical density.

The poor sinterability observed in monolithic GdB_4 is due to this material's high refractoriness and low intrinsic self diffusivity, which are also observed in other borides.^{13,14} Arabei *et al.*¹⁵ have reported that the densification pro-

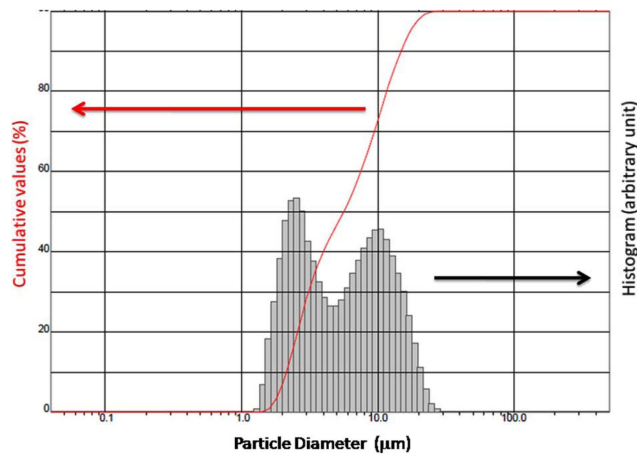


Fig. 3. Particle size distribution of GdB_4 powder.

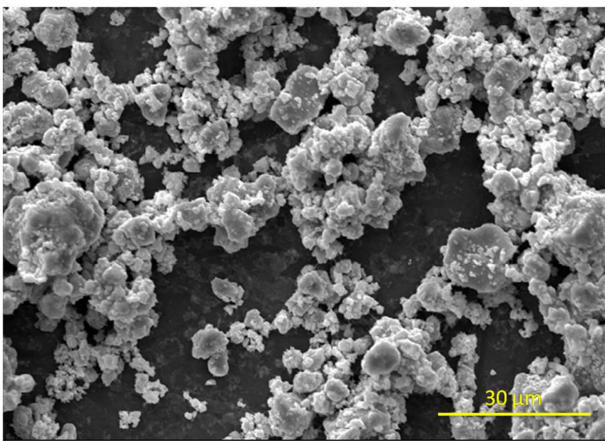
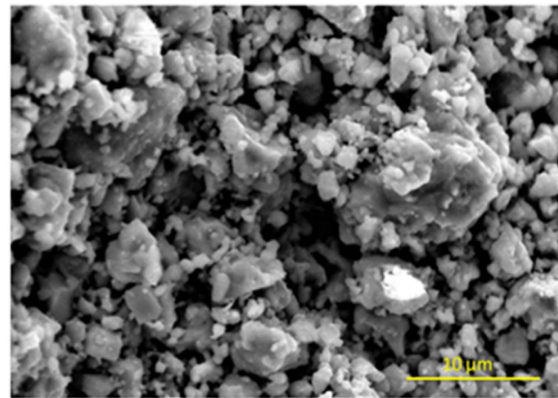


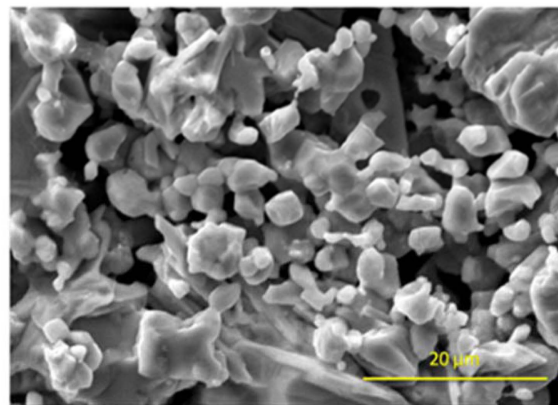
Fig. 4. SEM image of GdB_4 powder.

Table 4. Results of Densification of GdB_4

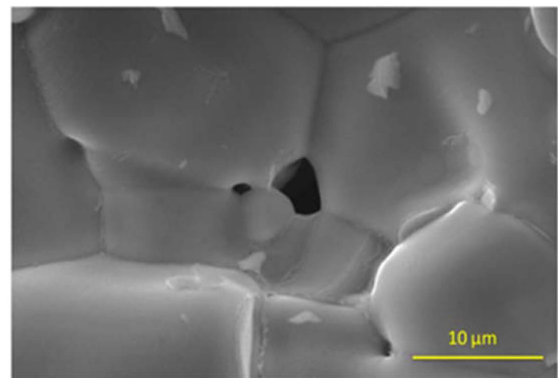
Technique	Temperature (°C)	Holding Time (h)	Relative Density (%)
Pressureless sintering	1200	2	63.3
	1400	2	65.2
	1600	2	65.5
	1800	2	77.1
Hot Pressing (35 MPa)	1700	2	95.5



(a)



(b)



(c)

Fig. 5. SEM images showing fracture surfaces of GdB_4 prepared by pressureless sintering at (a) 1200°C (b) 1400°C and (c) 1800°C.

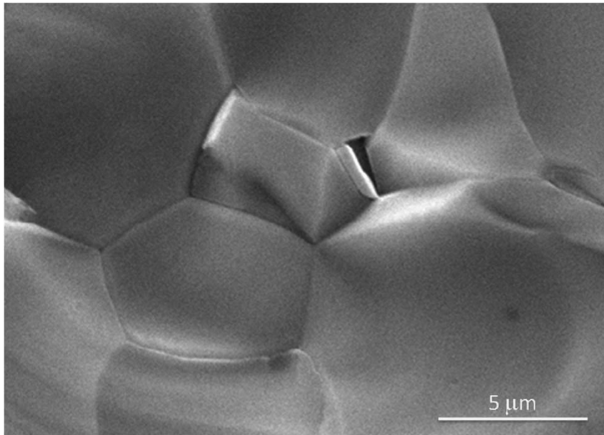


Fig. 6. Fracture surface of GdB₄ hot pressed at 1700°C and 35 MPa.

cesses of hexaborides becomes activated at temperatures in the range of 1800 - 2000°C. Zaykoski *et al.*¹⁷⁾ have prepared fully dense YB₄ by hot pressing at 1800°C.

Figure 5 provides an SEM image of the fracture surfaces of the pressureless sintered samples. Porosity is present in all the samples. The image also shows that there is no grain growth in the sample processed at 1200 and 1400°C. Grain growth was observed in the sample processed at 1800°C. Fig. 6 shows the fractured surface of the hot pressed GdB₄; this surface shows a dense microstructure.

3.3. Property characterization of hot pressed GdB₄

3.3.1. Mechanical Properties

The hardness of the dense GdB₄ sample (95.5% dense) was measured and found to be 21.4 GPa. In the literature, the hardness of GdB₄ was found to be 18.6 GPa.¹⁾ The fracture toughness of the dense GdB₄ sample (95.5% dense) was found to be 2.3 MPa·m^{1/2}. In the literature, there are no data available on fracture toughness measurement of GdB₄.

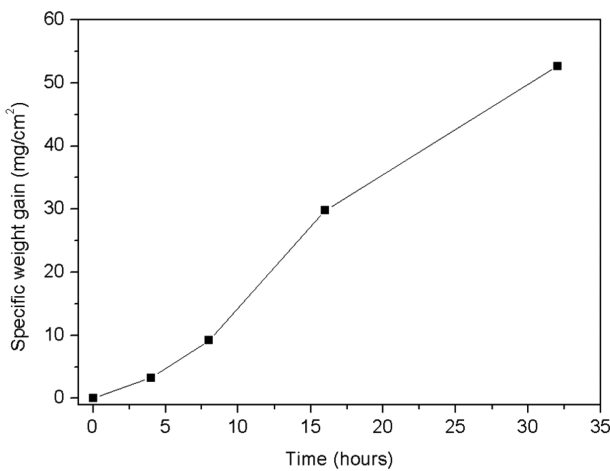
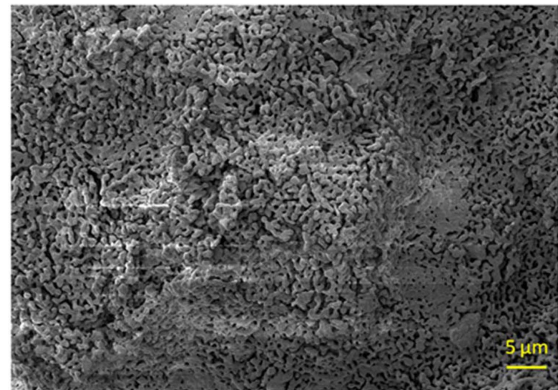


Fig. 7. Specific weight gain with time in GdB₄ sample during oxidation at 900°C (line joining points are for visual aid only).

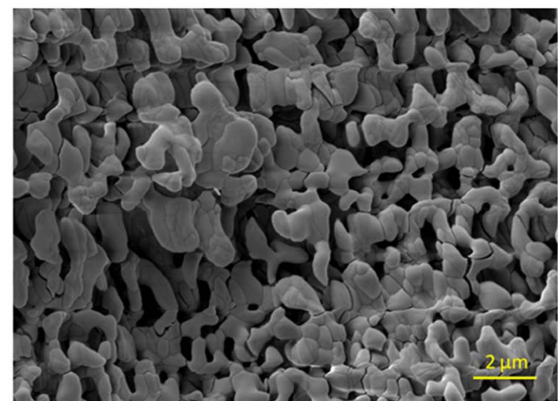
3.3.2. Oxidation study

Figure 7 presents the specific weight gain vs. time plot during isothermal oxidation of GdB₄ in air at 900°C. The plot shows linear behavior of the weight gain with time, which is a signature of a non-protective oxide layer. Specific weight gain values of 3.27 mg/cm² and 52.66 mg/cm² were measured after 4 h and 32 h, respectively.

Figure 8 provides an SEM image of the oxidized surface after oxidation. This image shows the presence of a porous



(a)



(b)

Fig. 8. SEM image of GdB₄ oxidized at 900°C for 32 h.

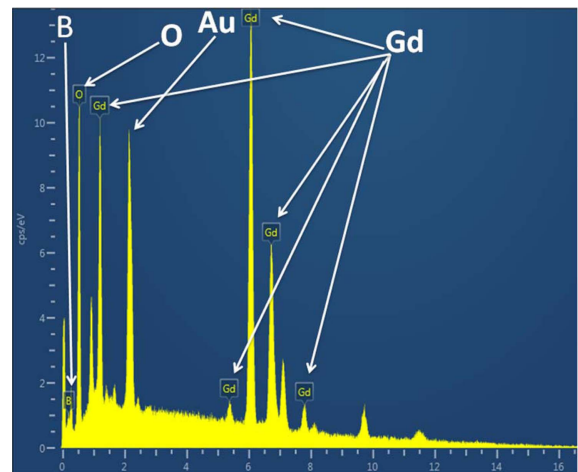


Fig. 9. Typical EDS pattern of oxidized surface of GdB₄.

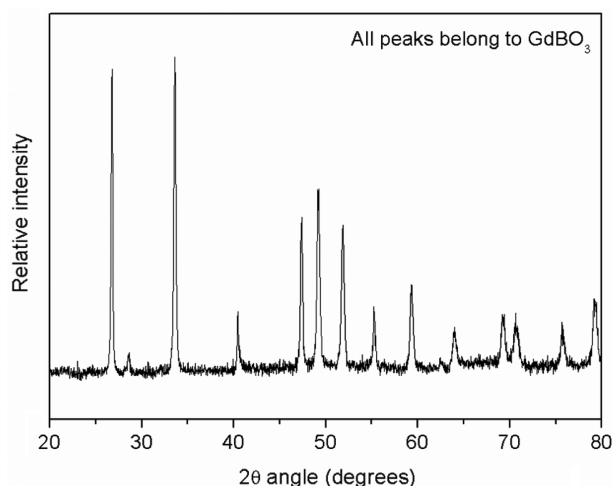
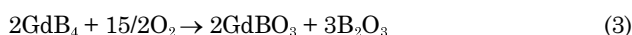


Fig. 10. XRD pattern of GdB_4 oxidized at $900^\circ C$ for 32 h.

and non-protective oxide layer on the surface. Fig. 9 shows the typical EDS pattern of the oxide layer. This indicates the presence of gadolinium, oxygen and boron on the oxidized surface. Fig. 10 shows the XRD pattern of the oxidized surface, which indicates the presence of $GdBO_3$ only as a crystalline phase. $GdBO_3$ could have been formed by reaction (3) during oxidation.



Free energy of formation for reaction (3) could not be calculated because thermodynamic data for GdB_4 are not available.

There has been no report on the oxidation behaviour of GdB_4 . In one study,¹⁷⁾ pure YB_4 was found to show considerable oxidation above $1200^\circ C$, and also resulted in the formation of a borate phase (YBO_3), similar to the oxidation behaviour of GdB_4 shown in the present study.

4. Conclusions

Conclusions that can be obtained from the present study are summarized as follows.

1. GdB_4 powder was synthesized by boron carbide reduction of Gd_2O_3 at $1500^\circ C$ in vacuum.
2. Pressureless sintering study of monolithic GdB_4 revealed that there is no densification until $1600^\circ C$.
3. It was possible to achieve a maximum density of 77.1% of the theoretical density at $1800^\circ C$.
4. Full densification was achieved by using hot pressing of GdB_4 powder at $1700^\circ C$.
5. Hardness and fracture toughness of fully dense GdB_4 was measured and found to be 21.4 GPa and $2.3 \text{ MPa}\cdot\text{m}^{1/2}$, respectively.
6. Oxidation study of monolithic GdB_4 revealed the formation of a non-protective, porous oxidized layer consisting of $GdBO_3$ after isothermal oxidation testing at $900^\circ C$ in an air atmosphere.

REFERENCES

1. M. L. Baucio, ASM Engineered Materials Reference Book; pp.300-9, ASM International, USA, 1994.
2. E. W. Hoyt, "Rare Earth Borides for Nuclear Applications"; pp.287-89 in *Proceedings of the Second Conference on Rare Earth Research*. Breach Science Publishers, New York, 1961.
3. P. Dunner, H. J. Heuvel, and M. Horle, "Absorber Materials for Control Rod Systems of Fast Breeder Reactors," *J. Nucl. Mater.*, **124** 185-94 (1984).
4. B. Weidenbaum, E. W. Hoyt, D. L. Zimmerman, W. V. Cummings, K. C. Antony, "Properties of Some High Temperature Control Materials," pp. 315-43 in *Materials and Fuels for High Temperature Nuclear Energy Applications*. Ed. by M. T. Simnad and M. R. Zumwalt, The M. I. T. Press, Massachusetts, 1962.
5. V. I. Lazorenko, A. P. Galasun, N. I. Siman, A. I. Dmitriev, and A. V. Brodovoi, "Preparation and Properties of GdB_4 Single Crystals," *Soviet Powder Metall. Met. Ceram.*, **29** [3] 224-27 (1990).
6. E. K. Storms and B. A. Mueller, "Thermionic Emission and Atom Vaporization of the Gd-B System," *J. Appl. Phys.*, **52** [4] 2966-70 (1981).
7. D. M. Goebel, Y. Hirooka, and T. A. Sketchley, "Large-Area Lanthanum Hexaboride Electron Emitter," *Rev. Sci. Instrum.*, **56** [9] 1717-22 (1985).
8. J. F. Rodríguez, J. A. Blanco, K. Katsumata, A. Kikkawa, F. Iga, and S. Michimura "Experimental Evidence of Non-collinear Magnetism in Gadolinium Tetraboride," *Phys. Rev. B*, **72** [5] 052407 (2005).
9. M. T. Garland, J. P. Wiff, J. Bauer, R. Guérin, and J. Y. Sailer, "The X-ray and Electronic Structures of GdB_4 ," *Solid State Sci.*, **5** [5] 705-10 (2003).
10. A. Baranovskiy and A. Grechnev, "Electronic Structure and Exchange Interactions in GdB_4 ," *J. Magn. Magn. Mater.*, **375** 96-9 (2015).
11. W. Wu, J. Xu, X. Bian, G. Hu, S. Sun, and G. Tu, "Phase Constitution in Mixed Gd_2O_3 and B_4C by Sintering at High Temperature," *J. Rare Earths*, **23** [3] 345-34 (2005).
12. H. Itoh, Y. Tsuzuki, T. Yogo, and S. Naka, "Synthesis of Cerium and Gadolinium Borides Using Boron Cage Compounds as a Boron Source," *Mater. Res. Bull.*, **22** [9] 1259-66 (1987).
13. J. K. Sonber, K. Sairam, T. S. R. C. Murthy, A. Nagraj, C. Subramanian, and R. C. Hubli, "Synthesis, Densification and Oxidation Study of Lanthanum Hexaboride," *J. Eur. Ceram. Soc.*, **34** [5] 1155-60 (2014).
14. J. K. Sonber and A. K. Suri, "Synthesis and Densification of ZrB_2 : Review," *Adv. Appl. Ceram.*, **110** [6] 321-34 (2011).
15. B. G. Arabei, E. N. Shtrom, and Y. A. Lapitskii, "Characteristics of the Manufacturing Technology of Dense Parts from, and the Mechanical Properties of, Some Hexaborides of the Rare-Earth Metals," *Powder Metall. Met. Ceram.*, **3** [5] 406-9 (1964).
16. G. R. Anstis, P. Chantikul, B. R. Lawn, and D. B. Marshall, "A Critical Evaluation of Indentation Techniques for Measuring Fracture Toughness: I, Direct Crack Measurements," *J. Am. Ceram. Soc.*, **64** [9] 533-38 (1981).
17. J. A. Zaykoski, M. M. Opeka, L. H. Smith, and I. G. Talmy,

- “Synthesis and Characterization of YB₄ Ceramics,” *J. Am. Ceram. Soc.*, **94** [11] 4059-65 (2011).
18. Y. Liu, W. J. Lu, J. N. Qin, and D. Zhang, “A New Route for the Synthesis of NdB₆ Powder from Nd₂O₃-B₄C System,” *J. Alloys Compd.*, **431** [1] 337-41 (2007).
19. X. H. Xu, H. N. Xiao, W. M. Guo, P. Z. Gao, and S. H. Peng, “Preparation and Reaction Mechanism of LaB₆ Powder by Solid-State Reaction at Atmospheric Pressure,” *J. Inorg. Mater.*, **26** 417-21 (2011).
20. J. K. Sonber, T. S. R. C. Murthy, C. Subramanian, R. C. Hubli, and A. K. Suri, “Synthesis, Densification and Characterization of EuB₆,” *Int. J. Refract. Met. Hard Mater.*, **38** 67-72 (2013).
21. J. K. Sonber, T. S. R. C. Murthy, K. Sairam, B. Paul, and J. K. Chakravartty, “Effect of TiSi₂ Addition on Densification of CeB₆,” *Ceram. Int.*, **42** 891-96 (2016).