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Unravelling the causes of the instability of $\text{Au}_n(\text{SR})_x$ nanoclusters on metal surfaces

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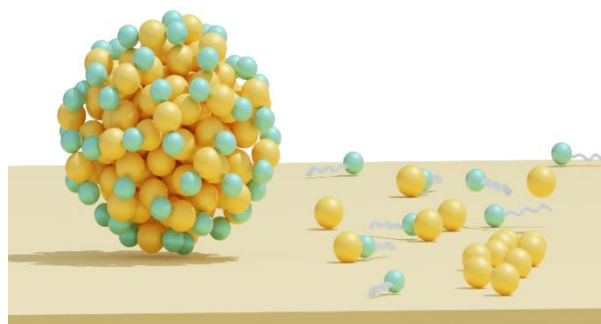
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ABSTRACT. Properties of small metal nanoclusters (NCs) rely on the exact arrangement of a few atoms. Minor structural changes can rapidly destabilize them, leading to disintegration. Here, we evaluate the energetic factors accounting for the stabilization and integrity of thiolated-capped AuNCs. We found that the core-cohesive and shell-binding energies regulate the disintegration process on a solid substrate by investigating the different energetic contributions, as shown here in a combined experimental and theoretical study. As the AuNC size increases, the core-cohesive energy and shell-stability (imposed by S-Au and hydrocarbon chain interactions) counterbalance the AuNC-substrate interaction and slow down the AuNC disintegration. Thus, the decomposition cannot only be understood in terms of desorption and transfer of the capping molecules to the support substrate, but conversely as a whole where ligand and core interactions play a role. Taken together, our experimental and theoretical results serve as guidelines for enhancing the stability of AuNC on solid-state devices, a key point for reliable nanotechnological applications as heterogeneous catalysis and sensing.

SYNOPSIS



INTRODUCTION

Understanding the physicochemical contributions that control nanocluster (NC) stability in different environments is crucial for using these attractive nanomaterials in technological applications.

There have been significant advances in the synthesis, structural, and physicochemical characterization of thiolate (SR)-protected gold nanoclusters (AuNC) in the last years. Today, the AuNC structure and stability are well described by the 'divide and protect' theory¹⁻² that explains the crystallographic data showing a metallic gold core protected by a shell of thiol gold complexes, $RS-(Au-SR)_x$, where x varies with the AuNCs size.³⁻⁸

Experimental and theoretical works show that the shell structure plays a critical role in its stability and reactivity when AuNCs are part of colloidal systems. In fact, simple RS ligand exchange can induce changes in size or structure, for instance, the transformation of $Ag_{25}(SR)_{18}$ into $Ag_{28}(SR)_{20}$ or $Ag_{144}(SR)_{60}$ into $Ag_{133}(SR)_{52}$.⁹ Small clusters such as $Au_{25}(SR)_{18}$ in solution can associate themselves to form dimers¹⁰ or form adducts with $Ag_{25}(SR)_{18}$ resulting in NC alloys.¹¹ These reactions involve opening the $RS-Au-(SR)-Au-SR$ shell, close sulfur-metal interactions between adjacent clusters and place exchange of metal atoms. The latter means a reactive $Au_{25}(SR)_{18}$ NC able to interact dynamically with other thiol-covered metallic clusters.

However, the 'divide and protect' theory does not explain the origin of the magic size NC stability. In this case, the superatom theory is employed.¹²⁻¹⁴ AuNCs are stable when their valence electrons make up a closed-shell electronic system. This theory accounts for most of the AuNC physicochemical properties but it is unable to predict the stability of gold nanoclusters universally.¹⁵ The thermodynamic stability theory proposed by Taylor and co-workers has recently

addressed this point.¹⁶ The theory predicts that the AuNC stability in colloid systems increases with the nanoparticle size, resulting in a delicate energy balance between the core cohesive energy and the shell-to-core binding energy. Moreover, it is predictive the links between the structure and, the stability and physico-chemical properties of AuNCs.¹⁷

Despite considerable progress in understanding the stability and reactivity of AuNC in colloid systems, relatively little attention has been paid to AuNC on solid supports. It has been shown that Au₂₀ deposited on ultrathin NaCl films exhibits sintering that results in a HOMO-LUMO gap reduction with increasing agglomerate size.¹⁸ Experimental works show that the interaction of AuNCs with solid-supports can destabilize the AuNCs leading, in the most extreme scenario, to the complete NC disintegration.¹⁹⁻²⁰ Recently, it has been found that both solid support and AuNC size have a strong influence on the heterogeneous catalysis of cyclohexane oxidation.²¹ The results show that, upon thermal air pretreatment, phenylethanethiolate-capped Au₁₄₄ clusters are more stable than their analogs Au₂₅ NCs. The higher stability was suggested to be related to the different cluster core structures, although it was pointed out that the core is stabilized by different staples motifs; Au₁₄₄ NC exhibits only short staples (SR-Au-SR) whereas Au₂₅ NC has long ones (SR-Au-SR-Au-SR), which may also influence the ligand removal upon pretreatment. Therefore, a key point for these nanomaterials with fascinating technological applications is to understand its behavior in contact with solid substrates.

Herein, we present an experimental and theoretical study of the stability of AuNCs on solid surfaces. Two benchmark AuNCs were chosen, Au₁₄₄(SR)₆₀ and Au₂₅(SR)₁₈ NCs. Results obtained for Au₁₄₄(SR)₆₀ placed in contact with clean reconstructed Au(111) were compared with those reported for Au₂₅(SR)₁₈ at the same NCs concentration and experimental conditions.^{19, 22} By looking up into the different energetic contributions, we found that the traditional picture where

the capping agent is the only one dictating the stability of nanoparticles does not apply in these systems. Instead, the core-cohesive energy and core-shell interactions regulate the process. These two factors counterbalance the nanocluster-substrate interaction hence, determining the degree of AuNC disintegration on the substrate.

EXPERIMENTAL SECTION

General. The glassware employed in this work was cleaned by immersion in boiling aqueous 20% HNO₃ solution for ~30 min. After this, the material was rinsed with ultrapure water and dried in an oven at 80 °C. For the AuNCs synthesis, an extra cleaning step with *aqua regia* solution was carried out to remove any metal traces. All reagents and solvents were purchased from Sigma-Aldrich and were used as received (purity as stated in the text). For aqueous solutions, ultrapure water (H₂O, 18.2 MΩ·cm, Purite Select Fusion 160, UK) was employed.

Au substrates. Two Au(111) substrates were employed. For STM measurements, we used Au(111) single crystals disks (99.999% purity, polished with roughness <0.01 μm and orientation accuracy <0.1°, MaTecK GmbH, Germany). For electrochemical experiments, we employed preferentially oriented Au(111) substrates (Arrandee™, Germany). Before use, the two substrates were cleaned by traditional methods, as stated in reference 19.

AuNCs synthesis and characterization. The hexanethiolate capped Au₁₄₄ and Au₂₅NCs were synthesized as previously reported^{19,23-24}, and a complete description of the method and characterization can be found in reference 19. In terms of unbonded thiolated species, the quality of purification was monitored by cyclic voltammetry and STM as follows. The AuNCs were dispersed in MeOH by vigorous shaking. After centrifugation (20 min, 13400 rpm), the supernatant was collected. Then, Au (111) substrates were immersed for 24 h in the supernatant,

rinsed with MeOH, and dried with N₂. Finally, the samples were studied by either cyclic voltammetry or STM. Voltammograms and STM images do not show any of the characteristic features ascribed to thiolates on Au, indicating that the unbonded-thiolates were largely removed after the purification steps performed.¹⁹

Multiple techniques confirmed the presence of the desired AuNCs as it was previously shown in reference 19, namely atomic force microscopy (AFM), scanning tunnelling microscopy (STM), differential pulse voltammetry and UV-Vis spectroscopy. UV-Vis spectra agree with uncharged AuNCs in both cases (see Figure S1).

Sample preparation. Au substrates were immersed in 2 μ M Au nanoclusters DCM dispersion, in the dark, and at room temperature, for the lapse of time indicated in the main text. Finally, they were rinsed with DCM and dried with N₂. For in-situ STM experiments in mesitylene, the Au substrates were imaged in mesitylene (98%) for about 10 min. Then, an aliquot of AuNC DCM dispersion was added to a final concentration of 3.4 μ M.

Electrochemistry. Electrochemical experiments were performed with a CHI760C potentiostat (CH Instruments, United States) and a conventional three-electrode glass cell. Pt coil and saturated calomel electrode (SCE) served as counter electrode (CE) and reference electrode (RE), respectively. Au substrate acted as the working electrode (WE). Cyclic voltammograms (CVs) were recorded in 0.1 M NaOH (99.99%) aqueous electrolyte at room temperature ($\sim 25^\circ\text{C}$). The electrolyte was degassed with Ar before the measurements, while an Ar atmosphere was maintained throughout the experiments.

Scanning Tunneling Microscopy (STM). STM measurements were performed using a Keysight 5100 STM (Keysight Technologies, United States). STM tips were prepared by mechanically

cutting a Pt/Ir wire (80:20%, 0.25 mm diameter, Goodfellow, UK). All images were acquired in constant-current mode using the tunneling conditions listed in Table S1. All images are shown with line-wise flattening to remove tilt in the substrate plane and a Gaussian filter to remove the noise. STM calibration was performed for each experiment by analyzing both the HOPG surface (x-y) and Au steps (z). STM images were analyzed with WSxM software (Nanotec Electronica S.L., Spain).²⁵

Calculations based on Density Functional Theory (DFT). The calculations were performed using the projector augmented wave (PAW) method,²⁶ as implemented in Vienna ab initio simulation package (VASP).²⁷⁻²⁹ The valence electrons were described within a plane-wave basis set and an energy cutoff of 420 eV; the remaining electrons were kept frozen as core states. Electron exchange-correlation was represented by the functional of Perdew, Burke, and Ernzerhof (PBE) of Generalized Gradient Approximation (GGA).³⁰ The weak van der Waals forces were treated by the semi-empirical method of Grimme (DFT-D) where the dispersion correction term is added to the conventional Kohn-Sham DFT energy³¹ with the parametrization DFT-D3.³² The energy convergence criterion was 10^{-5} eV for SCF energy except the calculation of Au₁₄₄(SR)₆₀ NC because the large numbers of atoms in the system (1344), the convergence process is slower than in the other calculations. The atomic positions were relaxed until the force on the unconstrained atoms was < 0.03 eVÅ⁻¹. The calculated Au lattice constant is 4.099 Å, which compares reasonably well with the experimental value (4.078 Å).³³ The Au₁₄₄(SR)₆₀ nanocluster consists of a hollow icosahedral Au₁₂ kernel. The second layer is composed of 42 gold atoms exhibiting 30 triangular (111) faces. The third layer is formed by 60 gold atoms which polyhedron surface consists of 12 pentagons and 20 equilateral triangles. This surface is wrapped by thirty RS-Au-SR staples. For all nanocluster structures calculated in this work, Au₁₄₄(SR)₆₀, Au₁₄₄, and

Au₁₄₄, the same cubic cell has been used of 70 Å × 70 Å × 70 Å where the nanocluster is separated by at least 35 Å of vacuum in all three directions. In these cases, the numerical integration was done only in the Γ point. All atomic positions of the adsorbates, including the gold atom in the staple complex in the NC, were allowed to relax in the optimization. On the other hand, the Au(111)-(1×1) substrate was represented by five atomic layers and a vacuum of ~17 Å that separates two successive slabs in our calculation. Surface relaxation was allowed in the three uppermost Au layers of the slab, while the atomic coordinates of the adsorbed species were allowed to relax without further constraints. Adsorbates were placed just on one side of the slab, and all calculations include a dipole correction. Optimal grid of Monkhorst-Pack.³⁴ k-points 3×7×1 were used for numerical integration in the reciprocal space of the (8×2√3) R30° unit cell. In the case of gas-phase, species was employed an orthogonal cell of appropriate size. Spin polarization was considered in all gas-phase species.

The SR and RS-Au-SR staple binding energies on Au₁₄₄(SR)₆₀ and (8×2√3) unit cell on Au(111) flat surface, E_b , are defined as follows,

$$(E_b^{SR})_{NC} = \frac{1}{60} [E(\text{Au}_{144}(\text{SR})_{60})_{NC} - E(\text{Au}_{144})_{NC} - 60E(\text{SR})_{\text{gas}}] \quad (1)$$

$$(E_b^{\text{staple}})_{NC} = \frac{1}{30} [E(\text{Au}_{144}(\text{SR})_{60})_{NC} - E(\text{Au}_{144})_{NC} - 30E(\text{RS-Au-SR})_{\text{gas}}] \quad (2)$$

$$(E_b^{SR})_{\text{Au}(111)} = \frac{1}{4} [E((\text{RS-Au-SR})_2)_{\text{Au}(111)} - E(\text{Au})_{\text{Au}(111)} - 4E(\text{RS})_{\text{gas}}] \quad (3)$$

where the NC/Au(111) subscript stands for nanocluster/Au(111) surface, and the SR/staple superscript refers to the species for which the binding energy is calculated. A negative number

indicates that adsorption is exothermic with respect to the separate clean surface and the adsorbate in the gas phase. On the other hand, the Gibbs free energy of adsorption of each surface structure (γ) can be approximated through the total energy from DFT calculations, the area (A), and the number of RS species (n) in the unit cell by using equation (4):

$$\gamma = \frac{n}{A} E_b \quad (4)$$

The cohesive gold energy has been calculated as:

$$E_c = E(\text{Au}_{\text{gas}}) - E(\text{Au}_{\text{bulk}}) \quad (5)$$

where $E(\text{Au}_{\text{gas}})$ represents the energy of the Au atom in the gas phase and $E(\text{Au}_{\text{bulk}})$ the energy of the NC complete (Au_{144}) or without the outer shell (Au_{114}) as appropriate.

RESULTS AND DISCUSSION

Figure 1 shows results illustrating the decomposition of $\text{Au}_{144}(\text{SR})_{60}$ on the Au(111) surface. STM images of the Au(111) substrate upon immersion in the nanocluster dispersion (12h, 2 μM in dichloromethane, DCM) reveal Au islands, monoatomic in height (0.24 nm, Figures 1a-b) with a surface coverage $\theta_{\text{Au_island}} = 0.23 \pm 0.02$ (Figure 1c), irrespective of the immersion time.

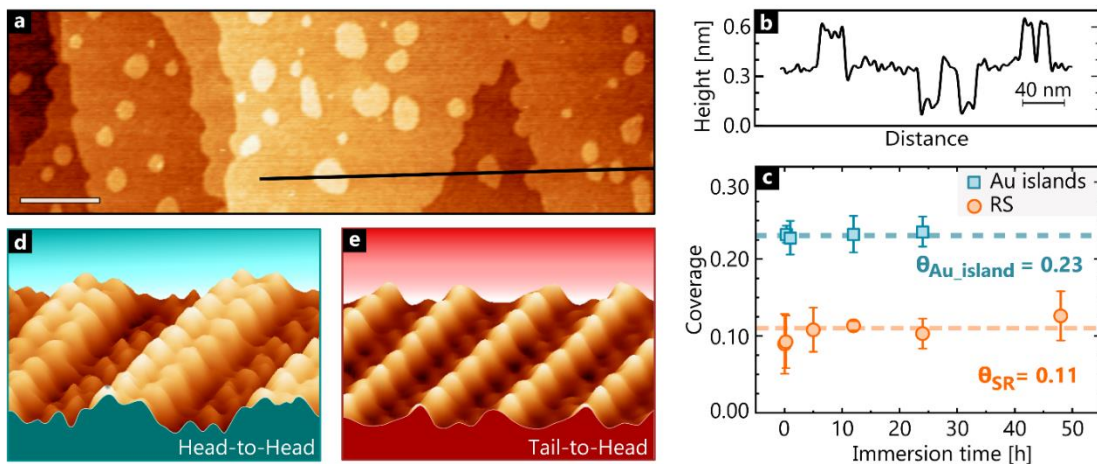


Figure 1. Characterization of Au(111) substrate upon immersion in a $Au_{144}(SR)_{60}$ dispersion in DCM. (a) STM images showing Au islands and thiol LD phases on the substrate surface. Scale bar corresponds to 60 nm. (b) Height profiles along the black line in image a. The profile shows that the Au islands are monoatomic in height (0.24 nm). (c) Temporal evolution of Au island and SR coverages. (d-e) High-resolution images of the two LD phases: Head-to-Head (d) and Head-to-Tail (e). The 3D images are $7 \times 3 \text{ nm}^2$ in size, with a rotational angle of 180° and a tilt of 17° .

The SR moieties arrange in two types of domains, as shown in Figures 1d-e. One domain corresponds to thiolated-species organized into a well-ordered phase of paired-rows separated by $2.3 \pm 0.1 \text{ nm}$. The distance between rows is $0.57 \pm 0.04 \text{ nm}$, and the bright spots forming each row are placed at $0.52 \pm 0.01 \text{ nm}$ (Figure 1d). This domain agrees with the head-to-head (H-H) lying-down phase (LD) reported previously.³⁵⁻³⁷ The second domain comprises rows of single bright spots separated by $1.26 \pm 0.02 \text{ nm}$ and $0.51 \pm 0.01 \text{ nm}$ (Figure 1e). This phase agrees with the head-to-tail (H-T) LD phase.^{35, 38-39}

On the other hand, the total number of thiol species determined by the charge of the thiolate-Au electroreduction peak (q_{SR}) results in a thiol coverage $\theta_{SR} = 0.11 \pm 0.01$ (Figure 1c), also irrespective of the immersion time. The θ_{SR} value and its time evolution (Figure 1c) suggest that

thiolated species are close-packed parallel to the Au(111) surface, reaching saturation in a relatively short time of immersion.

Before we move onto the system's energetic aspects, we will comment on the lack of vacancy islands (pits, Figure 1a). These features are fingerprints of alkanethiolate monolayers at high coverage ($\theta_{\text{SR}} = 1/3 \approx 0.33$).⁴⁰ At that coverage, molecules arrange in a dense standing-up phase (SU) as part of RS-Au-SR moieties on the Au(111) substrate. Au vacancies are proposed to emerge by removing of the Au adatoms needed to form the RS-Au-SR motif. By contrast, LD phases of hexanethiolate on Au(111) usually does not show Au vacancy islands.⁴¹ RS-Au-SR moieties can nevertheless still form, based on Au adatoms supplied through the lifting of the herringbone reconstruction, which provides a significant part of the required adatoms (0.043 out of 0.063).⁴²

For the LD phases observed here, two facts suggest the RS-Au-SR moieties' presence at the interface. First, the absence of the herringbone reconstruction that is present before immersion in the nanocluster (NC) dispersion (Figure S2) and second, the fact that the NC decomposition itself provides large amounts of Au adatoms. Indeed, by considering the experimental parameter θ_{SR} , we estimate that the expected Au island coverage would be 0.31 if AuNCs break apart into Au islands + SR, and 0.25 if break apart into Au islands + RS-Au-SR (cf. SI for further details). Comparison between experimental ($\theta_{\text{Au_island}} = 0.23$) and predicted values for the Au island coverage suggests that the lying-down structures contain adatoms as RS-Au-SR staple complex.

Thermodynamics of the thiols species on Au₁₄₄(SR)₆₀ and Au(111). To better understand the decomposition process of the NC Au₁₄₄(SR)₆₀, the electronic and geometric structure of the thiols species on the Au₁₄₄(SR)₆₀ and the Au(111) substrate have been investigated by DFT calculations.

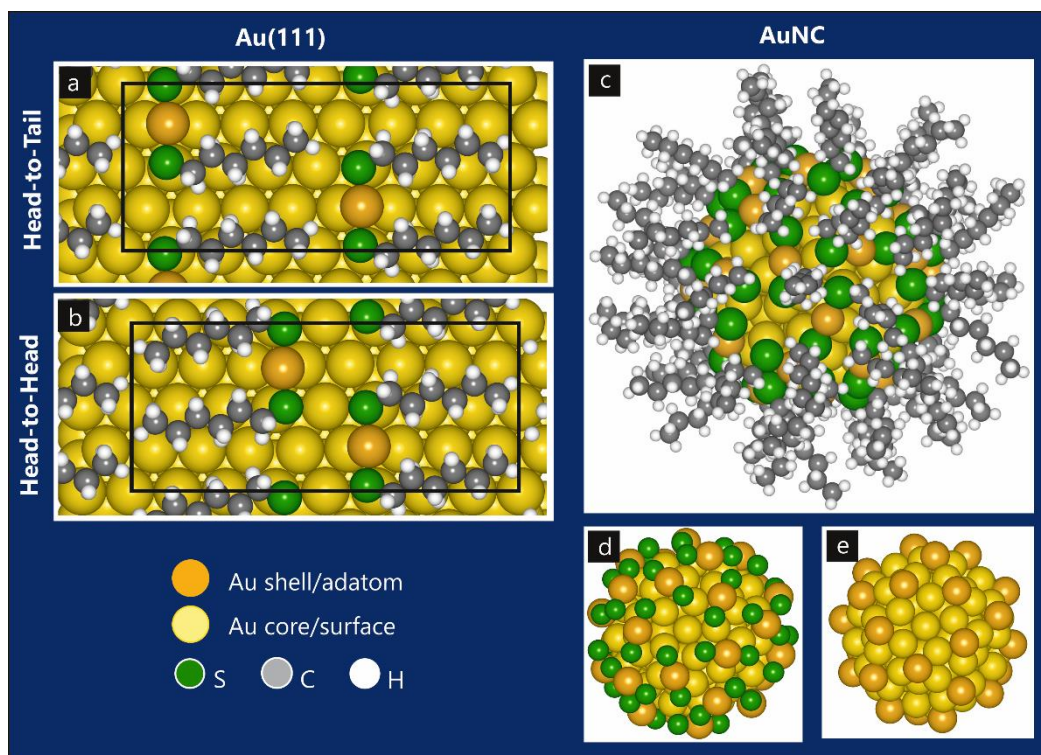


Figure 2. (a-b) Optimized structures for different LD domains on Au(111) : (a) RS-Au-SR in H-T, (b) RS-Au-SR on H-H. (c) Au₁₄₄(SR)₆₀ NC structure, (d) Au₁₄₄(SR)₆₀ NC without the alkyl chain, (e) Au₁₄₄ clean structure.

Our experimental data indicate that the AuNC protected by RS-Au-SR units decompose into Au islands and LD phases containing RS-Au-SR moieties upon interaction with the Au surface. We have recently shown that the most stable configurations for H-T and H-H phases are described, in both cases, by a $(8 \times 2\sqrt{3})_{\text{rect}}$ unit cell.⁴³ The optimized structures are shown in Figures 2a-b, respectively.

Table 1 shows the SR binding energy (E_b) and the surface free energy (γ) values obtained after the optimization of each surface structure (Figure 2). The E_b values of both LD models on the Au(111)

surface are higher than on the NC due to the optimization of the alkyl chain-alkyl chain and alkyl chain-substrate interactions.

Table 1. SR binding energy (E_b) and surface free energy (γ) for the different surface structures.

LD phase	Structures	Unit cell	θ_{RS}	E_b/RS [eV]	γ [meV \AA^{-2}]
H-T	[RS-Au-SR]@Au(111)	$(8 \times 2\sqrt{3})_{rec}$	1/8	-3.73	-64.08
H-H	[RS-Au-SR]@Au(111)	$(8 \times 2\sqrt{3})_{rec}$	1/8	-3.68	-63.22
---	[Au ₁₄₄ (SR) ₆₀] _{NC}	$(70 \times 70 \times 70)$	1	-2.85	-192.85

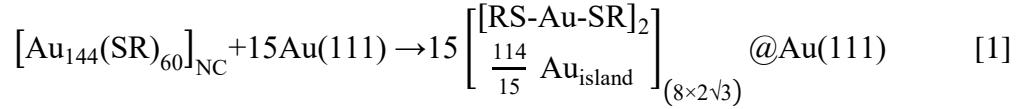
However, the thermodynamic stability is given by γ , as E_b is only an indication of how strong alkanethiolate species bond to the Au surface. Considering γ in the analysis, the SR radical on Au₁₄₄(SR)₆₀ is much more stable, by a factor of 3 compared to SR in LD lattices (Table 1). The highest stability of the SR species on the AuNC, despite its lower E_b , results from the more significant molecular density, and correspondingly, more considerable coverage ($\theta_{SR} = 1$) on the cluster than on the Au(111) surface (n/A term in equation 4). Thus, as with Au₂₅(SR)₁₈,²² the thermodynamic stability of the gold-adatom complexes by itself cannot explain the decomposition of the Au₁₄₄(SR)₆₀ NC.

The thermodynamic of the degradation process: small vs. large NCs. Au₁₄₄(SR)₆₀ and Au₂₅(SR)₁₈ degradation processes lead to the same products. The decomposition process results in Au islands and SR adsorbed on Au(111) as part of RS-Au-SR interfacial units for both nanoclusters.

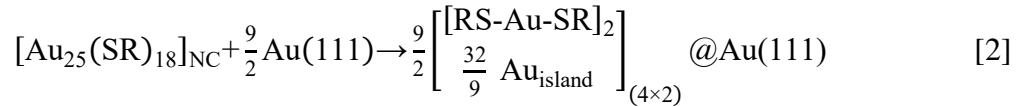
Despite the similarity, an intriguing observation is the different decomposition degrees observed between Au₂₅(SR)₁₈ and Au₁₄₄(SR)₆₀ NCs. Upon decomposition, Au₂₅(SR)₁₈ leads to an SU lattice

($\theta_{\text{SR}} = 1/3$). In contrast, for the same experimental conditions, $\text{Au}_{144}(\text{SR})_{60}$ only leads to a LD lattice ($\theta_{\text{SR}} = 1/8$) – note that this is the case even for extended immersion times (Figure 1c). The latter reflects that the number of decomposed $\text{Au}_{144}(\text{SR})_{60}$ is lower than that of $\text{Au}_{25}(\text{SR})_{18}$. Our estimation, based on the experimental θ_{SR} values, indicates that the number of decomposed $\text{Au}_{144}(\text{SR})_{60}$ is eight times lower than that of $\text{Au}_{25}(\text{SR})_{18}$ (cf. SI). This is unexpected considering that the initial AuNC number concentration is the same for both AuNCs, and more importantly, enough to fully saturate the Au(111) surface with the SR adopting the highest packing, i.e., molecules arranged in a $c(4 \times 2)$ lattice where SR adopts a SU configuration. Thus, it would be expected that both AuNCs decompose to the same degree.

Therefore, we evaluate the energetics of the complete decomposition process of the large single $\text{Au}_{144}(\text{SR})_{60}$ NC to form both H-T or H-H hexanethiolate lying-down phases according to reaction [1] and compare the result with that one obtained for the $\text{Au}_{25}(\text{SR})_{18}$ NC [2].²²



$$(\Delta E)_{\text{H-T}} = -68.5 \text{ eV and } (\Delta E)_{\text{H-H}} = -65.8 \text{ eV}$$



$$(\Delta E) = -21.9 \text{ eV}$$

The ΔE values show that the NC decomposition process is largely favored for a $\text{Au}_{144}(\text{SR})_{60}$ than for a $\text{Au}_{25}(\text{SR})_{18}$ NC, irrespective of the LD lattice. In fact, by setting the sight on the products, the $\text{Au}_{144}(\text{SR})_{60}$ decomposition process would be favored due to two factors. First, the stabilization provided by the Au island formation is higher for $\text{Au}_{144}(\text{SR})_{60}$ than for $\text{Au}_{25}(\text{SR})_{18}$, as the former

has a more available number of Au atoms per unit cell, 114 vs. 16. Second, the E_b value of SR species in the surface lattice is slightly higher for LD than for SU configuration (≈ -3.7 vs -3.5 eV, see Table 1).

The above argument is strictly valid at the zero-temperature limit. However, considering temperature through the entropic arguments would also favor the decomposition of the bigger AuNC. From the alkanethiolate point of view, the entropy change should be very little because their movement limitations for acquiring different conformations do not differ much between being adsorbed on a NC or a flat surface as Au(111). Instead, the process by which the core Au atoms are released from it to be part of Au islands adsorbed on the Au(111) surface implies an increase in the freedom degrees and thus an increase in entropy. This increase would be higher for Au₁₄₄ than Au₂₅ NCs because of the higher number of core atoms.⁴⁴ Therefore, if entropy would be considered in the energetic analysis, the conclusion would not change from that at $T = 0$ K.

To sum up, equations 1-2 indicate that the AuNC decomposition is favored for a Au₁₄₄(SR)₆₀ than for a Au₂₅(SR)₁₈. This conclusion contrasts the experimental data and points out that other factors should be considered in the stability analysis.

It is worth noting that reactions 1-2 only consider the initial and final state. In this sense, by looking into the difference within the AuNCs, we consider two potential factors that can hinder its decomposition: The Au cohesive energy and the outer shell composition.

The Au cohesive energy of the NCs is +3.00 eV for Au₁₄₄ and +2.43 eV for Au₂₅. However, if we consider the Au cohesive energy of these NCs without the Au external shell, the results hardly change for Au₁₁₄, +2.97 eV. In contrast, for the smallest NC, the Au₁₃ cohesive energy drops considerably, +2.02 eV (equation 5). In this sense, it is important to remember that the Au₂₅NC

has only one outer layer around the inner Au₁₃ kernel^{3, 8} while the Au₁₄₄ NC is made up of a Au₁₂ kernel surrounded by 3 layers of 42, 60, and 30 Au atoms, respectively.⁴⁵⁻⁴⁶

The second factor to be considered is the outer shell composition. For Au₁₄₄(SR)₆₀ NC, the shell is formed by thirty RS-Au-RS staple motifs that account for a total stabilization energy of -105.3 eV (each of the thirty complexes contributes with -3.51 eV, Table 2). In the case of Au₂₅(SR)₁₈, the stabilization energy imparted by the shell only accounts for -31.08 eV as the outer shell has only six RS-Au-(SR)-Au-SR staples (each of the six complexes contributes with -5.18 eV, Table 2). We also found that van der Waals energy increases with the cluster size and the number of staples in the shell. Table 2 shows that they contribute in 42% and 36% to the shell stability for Au₁₄₄(SR)₆₀ and Au₂₅(SR)₁₈, respectively. Taken together, the shell-composition also points out that the Au₁₄₄(SR)₆₀ presents a much higher shell-to-core interaction energy than the Au₂₅(SR)₁₈.

Table 2. Binding energy (E_b) and van der Waals energy (E_{vdW}) for the RS-Au-(SR)-Au-SR and SR-Au-SR complexes on AuNC's core.

Species	E_b [eV]	E_{vdW} [eV] ^[a]
[RS-Au-SR]@Au ₁₁₄	-3.51	-1.47 (42%)
[RS-Au-(SR)-Au-SR]@Au ₁₃	-5.18 ²²	-1.88 (36%)

[a] Values in parentheses correspond to the contribution of E_{vdW} to the E_b .

After this analysis, one can conclude that the larger cluster's improved stability on the Au(111) substrate results from both higher cohesive and shell-to-core interaction energies. These energetic contributions hinder to some extent the Au₁₄₄(SR)₆₀ decomposition process on the Au surface. Next, we will present experimental evidence that supports this conclusion by in situ imaging of the adsorption/decomposition process on the Au substrate.

The adsorption and decomposition of $\text{Au}_{144}(\text{SR})_{60}$ and $\text{Au}_{25}(\text{SR})_{18}$ on the Au(111) surface. In principle, for two particles capped with the same alkanethiol thiol, one would expect the larger particles' preferred adsorption *via* van der Waals interactions since they exhibit more contact sites with the surface.⁴⁷ Thus, the lower number of $\text{Au}_{144}(\text{SR})_{60}$ adsorbed and decomposed on the Au substrate cannot be explained by a difference in contact-sites between NCs and substrate surface. One could argue that the differences in size could affect either mobility in the DCM solution or introduce steric effects during adsorption and decomposition favoring the smaller NC. The number of $\text{Au}_{144}(\text{RS})_{60}$ that reached the surface is only 2.4 higher than $\text{Au}_{25}(\text{RS})_{18}$ if diffusion is considering; and even lower (1.5) if steric effect regulates the AuNC adsorption (cf. SI). Hence, none of these effects explains the experimental results.

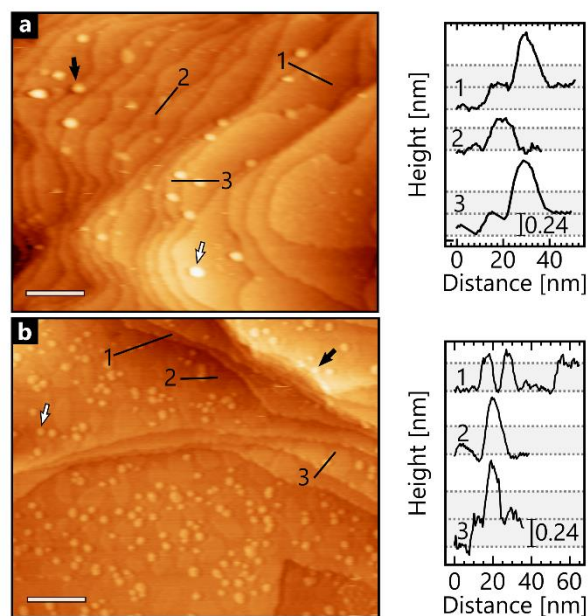


Figure 3. In situ STM images recorded in mesitylene upon addition of (a) $\text{Au}_{144}(\text{SR})_{60}$ and (b) $\text{Au}_{25}(\text{SR})_{18}$. The images show that for the big nanocluster the degradation products are preferentially located at the steps (black arrows), while those for $\text{Au}_{25}(\text{SR})_{18}$ on the terraces (white arrows). Height profiles are included in the left panels to highlight that bright spots correspond to degradation products rather than AuNCs. The dotted lines and the shadow areas within them indicate the step height in Au(111). Scale bars correspond to 60 nm.

A detailed analysis of the STM images in Figure 3 provides evidence of the adsorption/decomposition processes' factors. While the Au products originated by adsorption/decomposition of $\text{Au}_{144}(\text{SR})_{60}$ NCs are preferentially present at step edges (Figure 3a), those formed by the disintegration of $\text{Au}_{25}(\text{SR})_{18}$ NCs are randomly distributed on the Au(111) substrate (Figure 3b). It means that the $\text{Au}_{144}(\text{SR})_{60}$ shell, formed by RS-Au-SR moieties, only interacts upon contact with low-coordinated sites (more reactive) of the Au surface. On the other hand, the shell of the smaller $\text{Au}_{25}(\text{SR})_{18}$ NC formed by RS-Au-(SR)-Au-SR can also interact with the defective-free terrace (less reactive), as shown in Figure 3b. The reactivity of these small clusters is not surprising as they can form either $\text{Au}_{25}(\text{SR})_{18}$ dimers¹⁰ or adducts with $\text{Ag}_{25}(\text{SR})_{18}$ resulting in NC alloys.¹¹ These transformations involve the RS-Au-(SR)-Au-SR shell's spontaneous opening, S-metal close interactions between adjacent clusters, and place exchange of metal atoms. This means a reactive $\text{Au}_{25}(\text{SR})_{18}$ NC able to interact dynamically with thiol covered metal surfaces. This conclusion agrees with simulations about the dynamics of the core-shell bonds in Au_{25} clusters that suggest a shell loosely bound to the core.⁴⁸

On the other hand, it has been shown that NC reactivity decreases with the cluster size²¹. In this sense, a recent study on the chiral inversion of thiolated-protected gold NCs shows that the racemization mechanism and stability strongly depend on the size, core structure, and details of the metal-ligand structure.⁴⁹ Thus, a more rigid and less reactive $\text{Au}_{144}(\text{SR})_{60}$ NC is expected.

At this stage, some discussion about the solvent's role in opening the NC shell is needed. While gold island formation is a clearly surface-assisted process involving Au adatom surface diffusion, the formation of the alkanethiol phases could involve transferring SR species to the solvent and readsorption on the substrate surface. In this regard, it has been found that solvent-capping

interaction can influence the ligand conformation⁵⁰ or to be even stronger enough to break the Au-S bonds.⁵¹ It has been shown that dodecanethiolate-capped gold nanoparticles of 3-4 nm in size lost their ligand-molecules from the nanoparticles faces when immersed in tetrahydrofuran (THF). This process involves a selective rupture of the Au-S bond induced by the solvent environment, since this process is not observed in ethanolic solutions. In our case, we can discard the solvent-assisted desorption in solution, even though THF and DCM have similar polarities. In fact, at the same NC concentration, Au₁₄₄(SR)₆₀ should leave more thiolate species to the solution than the Au₂₅(SR)₁₈. Therefore, the larger NC should also produce dense SU phases rather than the diluted LD experimentally observed (Figure 1).

This observation confirms that NC decomposition occurs at the metal substrate surface as reported for AuNC on polycrystalline Ag²⁰ and Au(111).¹⁹ It can be argued, however, that the solvent would also play a key role in favoring the 'opening' of the RS-Au-SR or RS-Au-(SR)-Au-SR shell at the surface rather than breaking it in solution and, in this way, favoring transferring of Au-SR from the NC to the substrate. However, the effect of different solvents (ethanol, TFH, and DCM) on the Au-S stretching vibration frequency in a Au-SR complex performed by DFT calculations (not shown here) exhibit no differences concerning the value in the gas phase. In fact, it has been reported that the thermodynamic stability of AuNC is not affected by the solvents.¹⁶ Therefore, the S-Au breaking is a surface-assisted process where the NC shell's stability plays a key role and determines the number of thiolated species transferred from the NC to the Au(111) surface.

The mechanism of NC decomposition on Au(111). Based on the above observations, we propose the decomposition mechanism sketched in Figure 4. After adsorption on the Au(111) substrate, the NC shell is opened by interacting with the Au substrate atoms (Figure 4a).

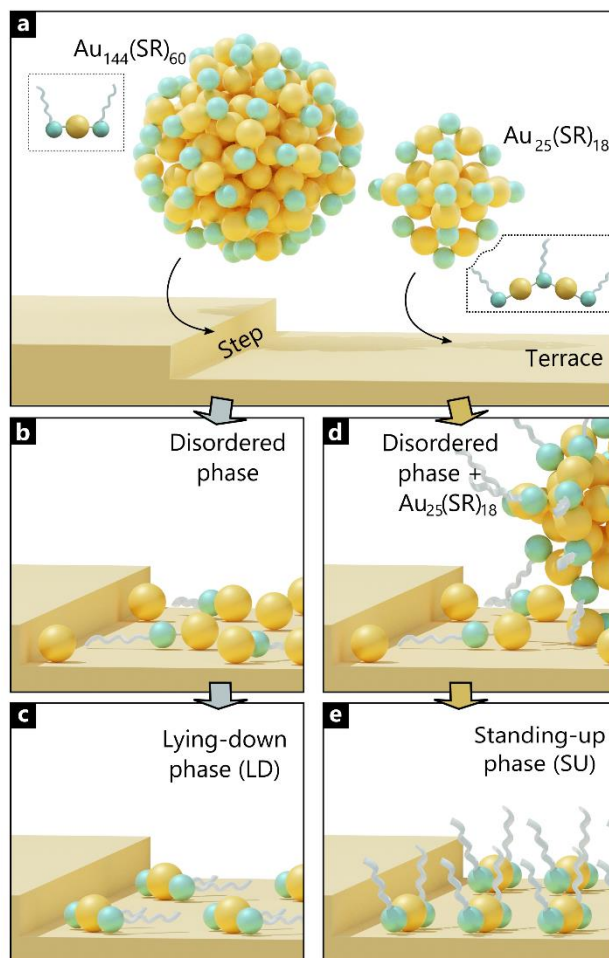


Figure 4. Sketch of the proposed decomposition process for $\text{Au}_{144}(\text{SR})_{60}$ (left panels) and $\text{Au}_{25}(\text{SR})_{18}$ (right panels). (a) The picture highlights the preferential adsorption site on the substrate: steps for $\text{Au}_{144}(\text{SR})_{60}$ and terraces for $\text{Au}_{25}(\text{SR})_{18}$. The insets show the structure of the staples forming the ligand-shell. Panels b and d illustrate the formation of a disordered phase resulting upon AuNCs break apart into their constituent units. As stated in panel d, $\text{Au}_{25}(\text{SR})_{18}$ adsorption is enhanced by the interaction between RS-Au-(SR)-Au-SR and the moieties already present at the substrate. (c-e) Thiolated monolayers obtained as a final product of the decomposition process: LD for $\text{Au}_{144}(\text{SR})_{60}$ (c) and SU for $\text{Au}_{25}(\text{SR})_{18}$ (e). For clarity, the Au islands are not included. Yellow: Au adatoms, light green: thiol radicals.

With the $\text{Au}_{144}(\text{SR})_{60}$, the ‘opening’ process occurs preferentially at step edges where the RS-Au-SR moieties can break into Au-SR and SR radicals that interact with the low coordinated Au atoms

(Figures 4a/b). In fact, recent in situ imaging of thiols adsorbed on Au(111) have shown that while RS-Au-SR moieties are present at terraces, SR radicals are present at step edges.⁵² The RS-Au-SR opening by AuNC interaction with the step edge leads to the NC shell's weakness and exposes the Au core atoms to the environment. Thus, AuNCs' components start to disintegrate, leading first to the Au-SR⁵³ and Au adatoms on the surface (Figures 4 b), and finally, to RS-Au-SR staples and Au islands. The fact that this process is restricted to the NC-step edge interaction besides with the high cohesive energy of this cluster result in a low population of RS-Au-SR, which accordingly accommodate in LD phases ($\theta_{\text{SR}} = 1/8$) and a relatively small number of Au islands (Figure 4c).

In contrast, for Au₂₅(SR)₁₈ NC, the opening of the RS-Au-(SR)-Au-SR shell takes place after interaction with any Au surface atom due to the high reactivity and dynamic of the shell following the mechanism already proposed in references ^{20, 22} (Figure 4d). Also, the adsorption and decomposition are enhanced by the interaction between RS-Au-(SR)-Au-SR and the moieties already present at the substrate (Au-SR or Au adatoms, Figure 4d). Later, the remaining AuNC's components easily disaggregate because of the low cohesion energy of the AuNC core and its interaction with the substrate surface. The surface reaction proceeds until the moieties present on the surface fully cover the surface and become unavailable for AuNC-surface dimerization, i.e., because of the high SR surface concentration, the thiolates adopt the standing-up configuration and become unavailable to be part of the 'bridge' complexes between surface-AuNC. This step is the formation of the SU phase ($\theta_{\text{SR}} = 1/3$) (Figure 4e).

CONCLUSIONS

We have discussed the stability of thiolated AuNCs on solids supports based on the degree of AuNC disintegration. We show experimental and theoretical results about the decomposition of

Au₁₄₄(SR)₆₀ nanoclusters on clean reconstructed Au(111) surfaces and compare them to those reported for Au₂₅(SR)₁₈ under the same experimental conditions.

Considering individual particles, the Au₁₄₄(SR)₆₀ decomposition is thermodynamically favored concerning the Au₂₅(SR)₁₈. However, its higher cohesive energy and shell stability hamper the NC disintegration to some extent. Therefore, Au₁₄₄(SR)₆₀ disintegration takes place preferentially at the high reactive site of the substrates where the stable shell can be opened - exemplified here with low-coordinated atoms of step edges. In contrast, Au₂₅(SR)₁₈ disintegration occurs over the entire substrate surface, confirming the flexibility and reactivity of the shell and the more straightforward core decomposition. Interestingly, our results suggest that the decomposition cannot be understood in terms of the desorption of individual thiol-gold adatom complexes since the RS-Au-(SR)-Au-SR dimeric staple has larger E_b than RS-Au-SR monomeric staple on their corresponding clusters but in terms of the rupture of the shell as a whole, where vdW forces play a significant role. Our experimental and DFT results confirm theoretical predictions about the ligand shell's increased stability with the increase in NC size proposed for colloid systems and the extent of its validity to solid supports.

Taken together, our results point out that the AuNC reactivity can be understood by the structure of the Au-SR complexes in the shell, but the AuNC stability requires considering the AuNC as a whole, where core and shell interactions play an essential role.

ASSOCIATED CONTENT

Supporting Information. Additional STM images. Imaging acquisition parameters for STM images. Determination of the RS coverage from electrochemical data. Additional calculations.

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